

Question number	Answer	Marks	Guidance
1	$\text{Ca} : \text{N} : \text{O} = 30.35/40.1 : 21.20/14.0 : 48.45/16.0$ $= 0.7569 : 1.5286 : 3.028$ Formula = CaN_2O_4	B1 B1	
2 (a)	$2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$	B1	
2 (b)	$4.50 \times 10^{-3} \text{ mol}$	B1	
2 (c)	$n(\text{KClO}_3) = 3.0 \times 10^{-3} \text{ mol}$ $M(\text{KClO}_3) = 122.6 \text{ g mol}^{-1}$ mass of $\text{KClO}_3 = 0.3678 \text{ g}$	B1 B1 B1	
3	$M(\text{ZnSO}_4) = 161.5 \text{ g mol}^{-1}$ $n(\text{ZnSO}_4) = 6.57 \times 10^{-3} \text{ mol}$ $n(\text{H}_2\text{O}) = (1.893 - 1.061)/18 = 4.62 \times 10^{-2} \text{ mol}$ $x = n(\text{H}_2\text{O})/n(\text{ZnSO}_4) = 7$ Formula = $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$	B1 B1 B1 B1	
4	$M(\text{Na}_2\text{CO}_3) = 106.0 \text{ g mol}^{-1}$ Actual $n(\text{Na}_2\text{CO}_3) = 0.0234 \text{ mol}$ Theoretical $n(\text{Na}_2\text{CO}_3) = 0.0234 \times 100/65$ $= 0.0360 \text{ mol}$ Theoretical $n(\text{NaHCO}_3) = 0.0720 \text{ mol}$ Mass of $\text{NaHCO}_3 = 0.0720 \times 84.0 = 6.05 \text{ g}$	B1 B1 B1 B1 B1	
5 (a)	$\text{Fe}_2\text{O}_3(\text{s}) + 3\text{CO}(\text{g}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{CO}_2(\text{g})$	B1	
5 (b)	$M(\text{Fe}_2\text{O}_3) = 159.6 \text{ g mol}^{-1}$ $n(\text{Fe}_2\text{O}_3) = (10\,000 \times 10^6)/159.6 = 6.266 \times 10^7 \text{ mol}$ $n(\text{Fe}) = 1.253 \times 10^8 \text{ mol} = 6.992 \times 10^5 \text{ g}$ (6992 tonnes)	B1 B1 B1	
6 (a)	$\text{Na} : \text{N} : \text{O} = 27.1/23.0 : 16.5/14.0 : 56.4/16.0$ $= 1.178 : 1.179 : 3.525$ Formula = NaNO_3	B1 B1	
6 (b)	$2\text{NaNO}_3(\text{s}) \rightarrow 2\text{NaNO}_2(\text{s}) + \text{O}_2(\text{g})$	B1	

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Table of Contents Ocr Chemistry Exam Question Answers

1. Understanding the eBook Ocr Chemistry Exam Question Answers
 - The Rise of Digital Reading Ocr Chemistry Exam Question Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry Exam Question Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry Exam Question Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry Exam Question Answers
 - Personalized Recommendations
 - Ocr Chemistry Exam Question Answers User Reviews and Ratings
 - Ocr Chemistry Exam Question Answers and Bestseller Lists
5. Accessing Ocr Chemistry Exam Question Answers Free and Paid eBooks
 - Ocr Chemistry Exam Question Answers Public Domain eBooks
 - Ocr Chemistry Exam Question Answers eBook Subscription Services
 - Ocr Chemistry Exam Question Answers Budget-Friendly Options
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- ePub, PDF, MOBI, and More
- Ocr Chemistry Exam Question Answers Compatibility with Devices
- Ocr Chemistry Exam Question Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Chemistry Exam Question Answers
 - Highlighting and Note-Taking Ocr Chemistry Exam Question Answers
 - Interactive Elements Ocr Chemistry Exam Question Answers
- 8. Staying Engaged with Ocr Chemistry Exam Question Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry Exam Question Answers
- 9. Balancing eBooks and Physical Books Ocr Chemistry Exam Question Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry Exam Question Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Chemistry Exam Question Answers
 - Setting Reading Goals Ocr Chemistry Exam Question Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Chemistry Exam Question Answers
 - Fact-Checking eBook Content of Ocr Chemistry Exam Question Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web at the conclusion of each standard you will find a few questions that test yor knowledge and serve as examples of the type of questions you may asked on the eoc perhaps the most useful part lies near the end of the document there you will find two practice exams each with an answer key at the end of the document

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