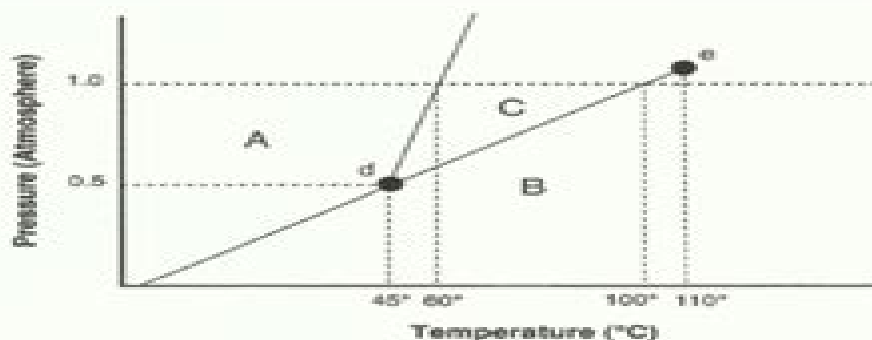


KEY

PHASE DIAGRAM WORKSHEET

Part A – Generic Phase Diagram.

Answer the questions below in relation to the following generic phase diagram.



1. Which section represents the solid phase? **A**
2. What section represents the liquid phase? **C**
3. What section represents the gas phase? **B**
4. What letter represents the triple point? **d**

In your own words, what is the definition of a triple point?

It is the point where all three phases coexist at a specific temperature and pressure

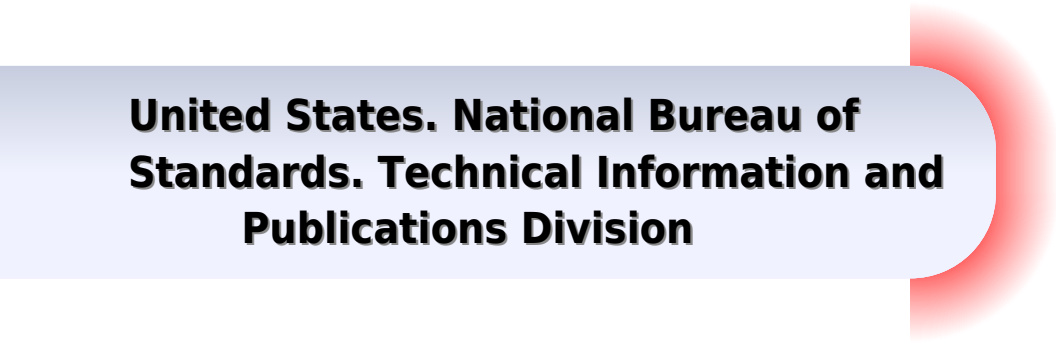
5. What is this substance's normal melting point, at 1 atmosphere of pressure? **60°C**
6. What is this substance's normal boiling point, at 1 atmosphere of pressure? **100°C**
7. Above what temperature is it impossible to liquefy this substance, no matter what the pressure?

110°C (This is known as the critical temperature)

8. At what temperature and pressure do all three phases coexist? **45°C and 0.5 atm**
9. At a constant temperature, what would you do to cause this substance to change from the liquid phase to the solid phase? **You would need to increase the pressure.**
10. What does sublimation mean? **The process of going from a solid to a gas directly, skipping the liquid phase altogether.**

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