



Name:

Date:

Student Exploration: Phase Changes (4b)

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: altitude, boil, boiling point, freeze, freezing point, gas, liquid, melt, melting point, phase, solid

Activity B:

Temperature and molecular motion

Get the Gizmo ready:

- Click **Reset**, and select the **Micro** view.
- Set **Ice volume** to 0 cm³.
- Set **Add/remove heat energy** to 0 J/s.



Question: Why do phase changes occur?

1. **Compare:** Set the **Water temperature** to 0 °C and click **Play**. Observe the water molecules. Click **Reset**, set the **Water temperature** to 100 °C, and click **Play** again.

What do you notice?

There is no difference.

2. **Observe:** Click **Reset**. The **mean molecular speed** of the water molecules is displayed below the container. Set the **Water temperature** to 0 °C and **Add/remove heat energy** to 400 J/s. Click **Play**.

A. How does the mean speed of the water molecules change as they are heated?

The water gets hotter (has more heat energy).

B. Does the mean molecular speed change as much as the temperature as the water heats up? Explain.

Yes, the more the temperature increases the more the molecules of the water speed up (until they hit the boiling point, where they cannot get hotter any longer.)

3. **Explain:** How is temperature related to the motions of molecules?

The higher the temperature, the faster the molecules move.

4. **Observe:** Click **Reset**. Set the **Water temperature** to 20 °C and the **Ice volume** to 50 cm³. Set **Add/remove heat energy** to 0 J/s. Click **Play**. How do the molecules in the liquid interact with the molecules in the solid?

Phase Changes Answer Key

**Thirumalaisamy P.
Velavancorresponding**



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Table of Contents Phase Changes Answer Key

1. Understanding the eBook Phase Changes Answer Key
 - The Rise of Digital Reading Phase Changes Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Phase Changes Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Phase Changes Answer Key
 - Personalized Recommendations
 - Phase Changes Answer Key User Reviews and Ratings
 - Phase Changes Answer Key and Bestseller Lists
5. Accessing Phase Changes Answer Key Free and Paid eBooks

- Phase Changes Answer Key Public Domain eBooks
- Phase Changes Answer Key eBook Subscription Services
- Phase Changes Answer Key Budget-Friendly Options
- 6. Navigating Phase Changes Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Phase Changes Answer Key Compatibility with Devices
 - Phase Changes Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Phase Changes Answer Key
 - Highlighting and Note-Taking Phase Changes Answer Key
 - Interactive Elements Phase Changes Answer Key
- 8. Staying Engaged with Phase Changes Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Phase Changes Answer Key
- 9. Balancing eBooks and Physical Books Phase Changes Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Phase Changes Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Phase Changes Answer Key
 - Setting Reading Goals Phase Changes Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Phase Changes Answer Key
 - Fact-Checking eBook Content of Phase Changes Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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