

pH Problems with Answers and Work Shown

pH Problems

Formulas:

$$\text{pH} = -\log[\text{H}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$[\text{H}^+] = 10^{-\text{pH}}$$

$$[\text{OH}^-] = 10^{-\text{pOH}}$$

$$[\text{H}^+][\text{OH}^-] = 10^{-14}$$

$$\text{pH} + \text{pOH} = 14$$

1. Calculate the pH of a 0.20 M HCl solution.
2. Calculate the pOH of a 0.0056 M solution of $[\text{OH}^-]$.
3. Calculate the $[\text{H}^+]$ of a solution with a pH of 5.
4. Calculate the $[\text{OH}^-]$ of a solution with a pOH of 4.75.
5. Calculate the $[\text{OH}^-]$ of a solution that has a pH of 8.
6. Calculate the $[\text{OH}^-]$ of a solution that that has a $[\text{H}^+] = 1.5 \times 10^{-4}$.



Ph Problems Answers

Chao Zhang



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