

If 16.0 g of  $\text{CaCO}_3$  are decomposed and 7.54 g of  $\text{CaO}$  are obtained, what is the percent yield of  $\text{CaO}$  obtained?  $\leftarrow$  A.Y



$$\text{Percent yield} = \frac{\text{Actual yield}}{\text{theoretical yield}} \times 100$$

$$16\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.1\text{g CaCO}_3} \times \frac{1\text{mol CaO}}{1\text{mol}}$$

# Percent Actual And Theoretical Yield Key

**Karin Nielsen-Saines**



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*Chemistry: A Basic Introduction* G. Tyler Miller, Jr., 1981      [Chemistry Textbook for College and University USA](#) Ibrahim Sikder, 2023-06-04

*Chemistry Textbook USA*      *AP Chemistry For Dummies* Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13

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