

CLASS 9 RS AGGARWAL SOLUTION



Rs Aggarwal Solution Class 9 Cbse

Richa Agarwal



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book allow pupils to feel the subject quite interesting Some distinctive features of this book are Strictly in accordance with the latest syllabus for the examinations to be held in and after 2025 Rational systematic and logical development of mathematical concepts All concepts and definitions have been discussed in detail and in a lucid manner The materials are class tested and the gradation has been done carefully The vocabulary and terminology used in the book is in accordance with the comprehension and maturity level of the students Subject matter is supplemented by suitable solved examples For each type of question set in exercises there are a number of solved examples Multiple Choice Questions have been added for better understanding of the lesson case Study Based Questions have been added to give a chance to the learners apply the mathematical knowledge gained to solve real life problems A well graded collection of questions with hints wherever required have been provided It is sincerely hoped that the present edition of the book would help the students to prepare adequately for the examinations It is also hoped that as the earlier books this book too will meet the needs and requirements of the teachers and students alike Suggestions for the improvement and notices of errors if any will be highly appreciated

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1
 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide naoh and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{NaOH} \cdot \text{Na}_2\text{C}_2\text{O}_4 \cdot 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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