

pH AND pOH

Name _____

The pH of a solution indicates how acidic or basic that solution is.

- pH range of 0 - 7 acidic
- 7 neutral
- 7-14 basic

Since $[H^+][OH^-] = 10^{-14}$ at 25°C , if $[H^+]$ is known, the $[OH^-]$ can be calculated and vice versa.

$$\text{pH} = -\log [H^+]$$

$$\text{So if } [H^+] = 10^{-6} \text{ M, pH} = 6.$$

$$\text{pOH} = -\log [OH^-]$$

$$\text{So if } [OH^-] = 10^{-8} \text{ M, pOH} = 8.$$

$$\text{Together, pH} + \text{pOH} = 14.$$

Complete the following chart.

	$[H^+]$	pH	$[OH^-]$	pOH	Acidic or Basic
1.	10^{-5} M	5	10^{-9} M	9	Acidic
2.		7			
3.			10^{-4} M		
4.	10^{-2} M				
5.				11	
6.		12			
7.			10^{-5} M		
8.	10^{-11} M				
9.				13	
10.		6			

Ph And Poh Chart Answer Key

George Tyler Miller



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