

## Experiment 25

### Advance Study Assignment: pH, Its Measurement and Applications

1. A solution of a weak acid was tested with the indicators used in this experiment. The colors observed were as follows:

|               |        |                   |        |
|---------------|--------|-------------------|--------|
| Methyl violet | violet | Congo red         | violet |
| Thymol blue   | orange | Bromocresol green | yellow |
| Methyl yellow | red    |                   |        |

What is the approximate pH of the solution?

\_\_\_\_\_

2. An aqueous solution of  $\text{NH}_3$  has a pH of 11.6. The ammonia ( $\text{NH}_3$ ) molecule is the conjugate base of the  $\text{NH}_4^+$  (called "ammonium") ion. Write the net ionic equation for the reaction that makes an aqueous solution of  $\text{NH}_3$  basic. (Eq. 5.)

3. The pH of a 0.10 M  $\text{HOBr}$  solution is 4.8.

- a. What is  $[\text{H}^+]$  in that solution?

\_\_\_\_\_ M

- b. What is  $[\text{OBr}^-]$ ? What is  $[\text{HOBr}]$ ? (Where do the  $\text{H}^+$  and  $\text{OBr}^-$  ions come from?)

\_\_\_\_\_ M; \_\_\_\_\_ M

- c. What is the value of  $K_a$  for  $\text{HOBr}$ ? What is the value of  $\text{p}K_a$ ?

\_\_\_\_\_

4. Formic acid,  $\text{HFor}$ , has a  $K_a$  value equal to about  $1.8 \times 10^{-4}$ . A student is asked to prepare a buffer having a pH of 3.90 from a solution of formic acid and a solution of sodium formate having the same molarity. How many milliliters of the  $\text{NaFor}$  solution should she add to 20. mL of the  $\text{HFor}$  solution to make the buffer? (See discussion of buffers.)

\_\_\_\_\_ mL

5. How many mL of 0.10 M  $\text{NaOH}$  should the student add to 20 mL 0.10 M  $\text{HFor}$  if she wished to prepare a buffer with a pH of 3.90, the same as in Problem 4?

\_\_\_\_\_ mL

# Ph Its Measurement Applications Advnaced Study

**A Loxley**



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