

The letters "pH" represent the French words "pouvoir hydrogene" which means "hydrogen power".

- The definition of pH is *pH is equal to the negative log (logarithm) of the hydrogen ion concentration of a solution.*
- The logarithm of a number is the power to which 10 must be raised to equal that number.

A pH value of less than 7 indicates a(n) acidic solution. A pH value of 7 indicates a neutral solution. A pH value of more than 7 indicates a(n) basic solution.

PROBLEMS: Show all work and circle the final answer.

1. Determine the pH of a 0.010 M HNO_3 solution.

$$\begin{aligned} \text{pH} &= -\log[0.010] \\ &= \boxed{2} \end{aligned}$$

2. What is the pH of a 2.5×10^{-6} M solution of HCl?

$$\begin{aligned} \text{pH} &= -\log[2.5 \times 10^{-6}] \\ &= \boxed{5.6} \end{aligned}$$

3. Calculate the pH of a solution of 0.0025M H_2SO_4 .

$$\begin{aligned} \text{pH} &= -\log[2 \times 0.0025] \\ &= -\log[0.005] \\ &= \boxed{2.3} \end{aligned}$$

4. Calculate the pH of a 0.0010 M NaOH solution.

$$\begin{aligned} [\text{H}^+][0.0010] &= 1.0 \times 10^{-14} \\ [\text{H}^+] &= \frac{1.0 \times 10^{-14}}{0.0010} \\ \text{pH} &= -\log[1 \times 10^{-11}] \\ &= \boxed{11} \end{aligned}$$

5. What is the pH of a 0.020M $\text{Sr}(\text{OH})_2$ solution?

$$\begin{aligned} [\text{OH}^-][0.04] &= 1 \times 10^{-14} \\ [\text{OH}^-] &= \frac{1 \times 10^{-14}}{0.04} \\ \text{pH} &= -\log[\text{H}^+] \\ &= -\log 2.5 \times 10^{-9} \end{aligned}$$

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Jacqueline Paris



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bioaccumulation and transformations in the atmosphere Structures chapters into basic and sophisticated sections Contains illustrative examples problems and case studies Examines the fundamental aspects of organic physical and inorganic chemistry applied to environmentally relevant problems Addresses problems and case studies in one volume Revise A2 Chemistry for Salters (OCR) Daniels,2005-04-04 Helps students to pull together key ideas in the course and apply them to exam questions in a fresh context Organised by module to allow readers to quickly access specific information this work provides tips on common pitfalls and advice on approaching exam questions with practice style exam questions for each module along with answers **Fundamentals of Chemistry: A Modern Introduction** Frank Brescia,2012-12-02

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