

Periodicity Chemistry Worksheet

A. Periodic table

1. Which are metals? Circle your answers: C, Na, F, Cs, Ba, Ni

Which metal in the list above has the most metallic character? Explain.

Cs - It is the metal with the ↑ shielding & ↓ ENC

2. Write the charge that each of the following atoms will have when it has a complete set of valence electrons forming an ion.

O⁻²

Na⁺¹

F⁻¹

N⁻³

Ca⁺²

Ar⁰

3. What is the most common oxidation number for calcium? Explain.

Ca⁺² - Calcium has 2 valence electrons, it will lose 2 electrons

Name two more elements with that oxidation number and explain your choice.

Any of the two: Be, Mg, Se, Ba, Ra - Same valence electrons / same family

4. What element in period 3 is a metalloid? Silicon

5. When element with atomic number 118 is discovered, what family will it be in? Noble Gas

Use the following word bank for questions 6 through 13

(Obviously, you will need to use several words more than once! ☹)

Alkali metals

Alkaline earth metals

Halogens

Noble gases

Transition metals

Noble gases

6. The Alkali metals have a single electron in the highest energy level.
7. The Alkaline earth metals achieve the electron configurations of noble gases by losing two electrons.
8. The transition metals are metals that can hold up to 10 electrons in their sublevel shape
9. The halogens achieve the electron configuration of noble gases by gaining one electron.
10. The noble gases have full s and p orbitals in the highest occupied energy levels.
11. The noble gases are stable and un-reactive.
12. The halogens are highly reactive and readily form salts with metals.
13. The Alkaline earth metals are metals that are more reactive than the transition elements but less reactive than the alkali metals.
14. Predict the oxidation number based on the electron configuration shown.
- | | | | |
|-----------------------|------------------------|-----------------------|------------------------|
| $1s^2 2s^2 2p^6 3s^2$ | <u>Mg⁺²</u> | $1s^2 2s^2 2p^6 3s^1$ | <u>Na⁺¹</u> |
| $1s^2 2s^2 2p^6$ | <u>Ne⁰</u> | $1s^2 2s^2 2p^5$ | <u>F⁻¹</u> |

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Table of Contents Periodic Worksheet Chemistry If8766 Answer

1. Understanding the eBook Periodic Worksheet Chemistry If8766 Answer
 - The Rise of Digital Reading Periodic Worksheet Chemistry If8766 Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Worksheet Chemistry If8766 Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Worksheet Chemistry If8766 Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Worksheet Chemistry If8766 Answer
 - Personalized Recommendations
 - Periodic Worksheet Chemistry If8766 Answer User Reviews and Ratings
 - Periodic Worksheet Chemistry If8766 Answer and Bestseller Lists
5. Accessing Periodic Worksheet Chemistry If8766 Answer Free and Paid eBooks

- Periodic Worksheet Chemistry If8766 Answer Public Domain eBooks
 - Periodic Worksheet Chemistry If8766 Answer eBook Subscription Services
 - Periodic Worksheet Chemistry If8766 Answer Budget-Friendly Options
6. Navigating Periodic Worksheet Chemistry If8766 Answer eBook Formats
- ePub, PDF, MOBI, and More
 - Periodic Worksheet Chemistry If8766 Answer Compatibility with Devices
 - Periodic Worksheet Chemistry If8766 Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Periodic Worksheet Chemistry If8766 Answer
 - Highlighting and Note-Taking Periodic Worksheet Chemistry If8766 Answer
 - Interactive Elements Periodic Worksheet Chemistry If8766 Answer
8. Staying Engaged with Periodic Worksheet Chemistry If8766 Answer
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Worksheet Chemistry If8766 Answer
9. Balancing eBooks and Physical Books Periodic Worksheet Chemistry If8766 Answer
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Worksheet Chemistry If8766 Answer
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Worksheet Chemistry If8766 Answer
- Setting Reading Goals Periodic Worksheet Chemistry If8766 Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Worksheet Chemistry If8766 Answer
- Fact-Checking eBook Content of Periodic Worksheet Chemistry If8766 Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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