

## Percent Yield Homework

I performed the following chemical reaction:



- 1) Balance the equation in the spaces above.
- 2) What type of reaction is taking place here? \_\_\_\_\_
- 3) Complete the equation by showing symbols of state as well as symbols around the arrow and H<sub>2</sub> values, if possible.
- 4) If I were to start with 40 grams of  $\text{C}_2\text{H}_5\text{OH}$  and an excess of oxygen, how many grams of carbon dioxide might I expect to make?
- 5) If my actual yield of carbon dioxide was 25 grams, what is my percent yield for this reaction?
- 6) Is my answer from problem #4 reasonable? If it is, explain why. If not, explain why not.
- 7) Explain why chemical reactions very rarely have a percent yield of 100%.

# Percent Yield Sheet

**United States. Department of  
Agriculture**



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