# Percent Composition and Molecular Formula Worksheet Key

- What's the empirical formula of a molecule containing 65.5% carbon, 5.5% hydrogen, and 29.0% oxygen? C<sub>3</sub>H<sub>3</sub>O mass = 55 g/mole
- If the molar mass of the compound in problem 1 is 110 grams/mole, what's the molecular formula?
   C<sub>6</sub>H<sub>6</sub>O<sub>2</sub>
- What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen? Li<sub>2</sub>CO<sub>3</sub>
- If the molar mass of the compound in problem 3 is 73.8 grams/mole, what's the molecular formula?
   Li<sub>2</sub>CO<sub>3</sub>

#### Write the molecular formulas of the following compounds:

- A compound with an empirical formula of C<sub>2</sub>OH<sub>4</sub> and a molar mass of 88 grams per mole. C<sub>4</sub>O<sub>2</sub>H<sub>8</sub>
- 6. A compound with an empirical formula of C4H4O and a molar mass of 136 grams per mole. C8H8O2
- A compound with an empirical formula of CFBrO and a molar mass of 254.7 grams per mole. C<sub>2</sub>F<sub>2</sub>Br<sub>2</sub>O<sub>2</sub>
- 8. A compound with an empirical formula of C2H8N and a molar mass of 46 grams per mole. C2H8N

#### Answer the following questions:

- The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid. CH<sub>2</sub>O
- 10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula? C<sub>2</sub>H<sub>4</sub>O<sub>2</sub>
- 11. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO<sub>2</sub>, H<sub>2</sub>O, and N<sub>2</sub> as products. If the combustion of 9.71 g of aniline yields 6.63 g H<sub>2</sub>O and 1.46 g N<sub>2</sub>, what is its empirical formula? C<sub>6</sub>H<sub>7</sub>N
- 12. The molar mass of aniline is 93 g/mol. What is its molecular formula? CoH1N
- Calculate the mass percent of carbon, nitrogen and oxygen in acetamide, C<sub>2</sub>H<sub>5</sub>NO. %C 40.668 %H 8.533 %N 23.713 %O 27.086
- 14. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound? PCIs
- When 2.5000 g of an oxide of mercury, (Hg<sub>x</sub>O<sub>y</sub>) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula. Hg<sub>2</sub>O
- 16. The compound benzamide has the following percent composition. What is the empirical formula?
  C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 % C<sub>7</sub>H<sub>7</sub>NO
- 17. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?

C<sub>3</sub>H<sub>2</sub>NO<sub>3</sub> empirical formula C<sub>4</sub>H<sub>2</sub>NO<sub>3</sub> molecular formula

# **Percent Composition And Molecular Formula Answers**

**Rosina Ehmann** 

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- 1. Understanding the eBook Percent Composition And Molecular Formula Answers
  - The Rise of Digital Reading Percent Composition And Molecular Formula Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Percent Composition And Molecular Formula Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Percent Composition And Molecular Formula Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Percent Composition And Molecular Formula Answers
  - Personalized Recommendations
  - Percent Composition And Molecular Formula Answers User Reviews and Ratings
  - Percent Composition And Molecular Formula Answers and Bestseller Lists

- 5. Accessing Percent Composition And Molecular Formula Answers Free and Paid eBooks
  - Percent Composition And Molecular Formula Answers Public Domain eBooks
  - Percent Composition And Molecular Formula Answers eBook Subscription Services
  - Percent Composition And Molecular Formula Answers Budget-Friendly Options
- 6. Navigating Percent Composition And Molecular Formula Answers eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Percent Composition And Molecular Formula Answers Compatibility with Devices
  - Percent Composition And Molecular Formula Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Percent Composition And Molecular Formula Answers
  - Highlighting and Note-Taking Percent Composition And Molecular Formula Answers
  - Interactive Elements Percent Composition And Molecular Formula Answers
- 8. Staying Engaged with Percent Composition And Molecular Formula Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Percent Composition And Molecular Formula Answers
- 9. Balancing eBooks and Physical Books Percent Composition And Molecular Formula Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Percent Composition And Molecular Formula Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Percent Composition And Molecular Formula Answers
  - Setting Reading Goals Percent Composition And Molecular Formula Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Percent Composition And Molecular Formula Answers
  - Fact-Checking eBook Content of Percent Composition And Molecular Formula Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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