

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
Calculate: Moles of product based on the limiting reagent. Then convert moles of product to grams. Percent yield = (actual yield / theoretical yield) x 100.

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
 Calculate: Moles of product based on the limiting reagent. Then convert moles of product to grams.



What mass of product will be produced from 10.0 g of Al?

Al	10.0 g
FeCl ₃	10.0 g
AlCl ₃	10.0 g
Fe	10.0 g

- 10.0 g of Al is the limiting reagent. 10.0 g of Al will produce 10.0 g of Fe. 10.0 g of Fe is the limiting reagent. 10.0 g of Fe will produce 10.0 g of Fe.



- 10.0 g of Al is the limiting reagent. 10.0 g of Al will produce 10.0 g of Fe. 10.0 g of Fe is the limiting reagent. 10.0 g of Fe will produce 10.0 g of Fe.

- 10.0 g of FeCl₃ is the limiting reagent. 10.0 g of FeCl₃ will produce 10.0 g of Fe.



- 10.0 g of FeCl₃ is the limiting reagent. 10.0 g of FeCl₃ will produce 10.0 g of Fe. 10.0 g of Fe is the limiting reagent. 10.0 g of Fe will produce 10.0 g of Fe.



Percent Yield And Limiting Reactants Quiz

Julia Schneider



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