

Table K
Selected Polyatomic Ions

Formula	Name	Formula	Name
ClO_4^-	perchlorate	ClO_2^-	chlorite
ClO_3^-	chlorate	ClO_2^+	dichlorine oxide
NH_4^+	ammonium	MnO_4^-	permanganate
$\text{C}_2\text{H}_3\text{O}_2^-$ CH_3COO^-	acetate	NO_2^-	nitrite
CN^-	cyanide	NO_3^-	nitrate
CO_3^{2-}	carbonate	O_2^{2-}	peroxide
HCO_3^-	bicarbonate hydrogencarbonate	OH^-	hydroxide
$\text{C}_2\text{O}_4^{2-}$	oxalate	PO_4^{3-}	phosphate
CrO_4^{2-}	chromate	HCO_3^-	hydrogencarbonate
CrO_2^-	chromite	SO_3^{2-}	sulfite
CrO_4^{2-}	chromate	SO_3^{2-}	sulfite
CrO_2^-	chromite	HSO_4^-	hydrogensulfate
CrO_4^{2-}	chromate	$\text{S}_2\text{O}_3^{2-}$	thiosulfate

Table L
Solubility Guidelines for Aqueous Solutions

Basic Most Forms Soluble Compounds	Exceptions	Basic Most Forms Soluble Compounds*	Exceptions
Group 1 ions (Li^+ , Na^+ , etc.)		acetate (CH_3CO_2^-)	when combined with Cu^{2+} or Pb^{2+}
ammonium (NH_4^+)		chromate (CrO_4^{2-})	when combined with Group 1 ions, Ca^{2+} , Sr^{2+} , or ammonium (NH_4^+)
nitrate (NO_3^-)		phosphate (PO_4^{3-})	when combined with Group 1 ions or ammonium (NH_4^+)
hydrogencarbonate (HCO_3^-)		oxide (O^{2-})	when combined with Group 1 ions or ammonium (NH_4^+)
halides (Cl^- , Br^- , I^-)	when combined with Ag^+ , Pb^{2+} , or Hg_2^{2+}	hydroxide (OH^-)	when combined with Group 1 ions, Ca^{2+} , Sr^{2+} , Ba^{2+} , or ammonium (NH_4^+)
sulfates (SO_4^{2-})	when combined with Ag^+ , Ca^{2+} , Sr^{2+} , Ba^{2+} , or Pb^{2+}		

*Transition metals generally have low solubility as Hg_2^{2+}

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