

Unit 3 Review - Atomic Structure & Periodic Table

Part 1: Answer the following questions.

1. a. Which scientist created the first modern atomic theory? John Dalton
 b. What was his theory?
 1. All matter is made of extremely small particles called atoms.
 2. All atoms of a given element are identical and atoms of a specific element are different from any other element
 3. Atoms cannot be created, divided into smaller particles or destroyed
 4. Different atoms combine in simple whole number ratios to form compounds
 5. In a chemical reaction, atoms are separated, combined or rearranged.
- c. What was his model of the atom called? Billiard Ball
- d. What was proven wrong and why?
 1. All atoms of a given element are identical – wrong b/c of the existence of isotopes
 2. Atoms cannot be divided into smaller particles – wrong b/c of the existence of electrons, protons and neutrons
2. a. Which scientist did the gold foil experiment? Rutherford
 b. What did he conclude from the gold foil experiment?
 1. Atoms are made up of mostly empty space
 2. Atoms have a small positive center call the nucleus
3. a. Who discovered the electron? J. J. Thomson
 b. What was the name of his experiment? Cathode Ray Tube Experiment
 c. What was his model of the atom called? Plum Pudding Model
4. What did Mendeleev do?
Arranged the periodic table in order of increasing atomic mass
5. a. Who arranged the **modern** periodic table? Moseley
 b. How is the modern periodic table arranged? Arranged in order of Atomic Number
6. a. What are the three subatomic particles? Give the symbols, charge, mass and location of each.

Particles	Symbol	Charge	Mass	Location
Proton	p^+	Positive	1 amu	Nucleus
Neutron	n^0	Neutral	1 amu	Nucleus
Electron	e^-	Negative	1/2000 amu	Electron Cloud

7. What is special about the proton? **The proton IDENTIFIES THE ELEMENT**
8. Fill in the blanks in the following:
 - a. Mass # = $p^+ + n^0$
 - b. Atomic # = # of protons
9. What is the difference between Average Atomic mass and mass #? (Average) Atomic Mass is found on the periodic table and is the weighted average of the isotopes. The Mass Number is the mass of one element and must be given or calculated by adding the protons and neutron of that element.
10. Which of the following are isotopes of the same element? How do you know? **A. and D. They have the same Atomic Number (# of Protons)**
 - a. $^{12}_5X$
 - b. $^{15}_5X$
 - c. $^{12}_6X$
 - d. $^{15}_5X$
11. Calculate the average atomic mass for Boron if it has two naturally occurring isotopes: boron -10 (abundance of 19.8%) and boron - 11 (abundance of 80.2%).

$$\left(10 \times \frac{19.8}{100} \right) + \left(11 \times \frac{80.2}{100} \right) = 10.802 \text{ amu}$$
12. Identify an element "X" by calculating the average atomic mass. X = 6 (abundance of 7.5%) and X= 7 (abundance 92.5%).

$$\left(6 \times \frac{7.5}{100} \right) + \left(7 \times \frac{92.5}{100} \right) = 6.925 \text{ amu} \quad \text{Which is Lithium}$$

Physical Science Atomic Structure Test Answers

S Ashworth



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Table of Contents Physical Science Atomic Structure Test Answers

1. Understanding the eBook Physical Science Atomic Structure Test Answers
 - The Rise of Digital Reading Physical Science Atomic Structure Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Atomic Structure Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Atomic Structure Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Atomic Structure Test Answers
 - Personalized Recommendations
 - Physical Science Atomic Structure Test Answers User Reviews and Ratings
 - Physical Science Atomic Structure Test Answers and Bestseller Lists

5. Accessing Physical Science Atomic Structure Test Answers Free and Paid eBooks
 - Physical Science Atomic Structure Test Answers Public Domain eBooks
 - Physical Science Atomic Structure Test Answers eBook Subscription Services
 - Physical Science Atomic Structure Test Answers Budget-Friendly Options
6. Navigating Physical Science Atomic Structure Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Atomic Structure Test Answers Compatibility with Devices
 - Physical Science Atomic Structure Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Atomic Structure Test Answers
 - Highlighting and Note-Taking Physical Science Atomic Structure Test Answers
 - Interactive Elements Physical Science Atomic Structure Test Answers
8. Staying Engaged with Physical Science Atomic Structure Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Atomic Structure Test Answers
9. Balancing eBooks and Physical Books Physical Science Atomic Structure Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Atomic Structure Test Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Atomic Structure Test Answers
 - Setting Reading Goals Physical Science Atomic Structure Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Atomic Structure Test Answers
 - Fact-Checking eBook Content of Physical Science Atomic Structure Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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