

Organic Compounds: carbon containing chains

Alkane- single bonds

Alkene- double bonds

Alkyne- triple bonds

Naming uses prefixes based on # carbons, then the root above

Depending on bond type

1 carbon = meth-

2 carbon = eth-

3 carbon = prop-

4 carbon = but-

5 carbon = pent-

6 carbon = hex-

7 carbon = hept-

8 carbon = oct-

9 carbon = non-

10 carbon = dec-

Monomers build up to be polymers

Nuclear Radiation: 1) elements above atomic # 83.

2) atoms with uneven #p and #n ratio

Nuclear splitting

Nuclear joining

Decay particles:

Alpha- α / 4_2 He, positive, most massive, like helium ion

Beta- β / $^-$ e, negative, neglig. Mass, like electron

Gamma- γ / 0_0 , no charge, no mass, most energy, most penetrating

Half life time for half to decay

100 gram sample has half life of 10 days. How much remains after 30

days?

30/10 = 3 half lives: $100\text{g}/2 = 50\text{g}/2 = 25\text{g}/2 = 12.5\text{g}$, grams remain

Chemical Reactions: new substance created.

Chemical Equations

Reactant $\rightarrow$ Product

Synthesis: $A + B \rightarrow AB$

Decomposition: $AB \rightarrow A + B$

Single Displacement: $A + BC \rightarrow AC + B$

Double Displacement: $AB + CD \rightarrow AD + CB$

Balanced Chemical equations: Law of Conservation of Mass

Use coefficients- large whole numbers put front of substance

$x\text{H}_2 + \text{O}_2 \rightarrow x\text{H}_2\text{O}$

Coefficients also indicate # moles of each substance taking part in the reaction & mole ratio

Ex: 2 mol H_2 react with 1 mol O_2 to produce 2 mol H_2O

Molar mass: mass of 1 mole of a substance.

Add up masses from PT for each ATOM present in substance.

Ex: H_2O : [2 atoms of H @ 1g each] + [1 atom of oxygen @ 16g each] =

18 g/mol

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