

Chapter 7: Chemical Reaction Guided Notes

I. Chemical Changes in Matter

A. Chemical Reaction

1. A change in which one or more substances are converted to _____ substances.

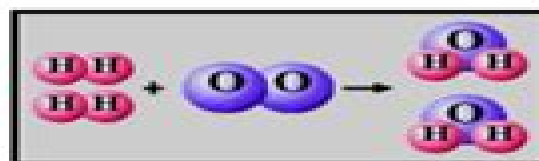


B. Law of Conservation of Matter

1. In a chemical reaction, matter is not _____ or _____.
2. Atoms can only be _____.
3. Discovered by Lavoisier.

_____ Hydrogen

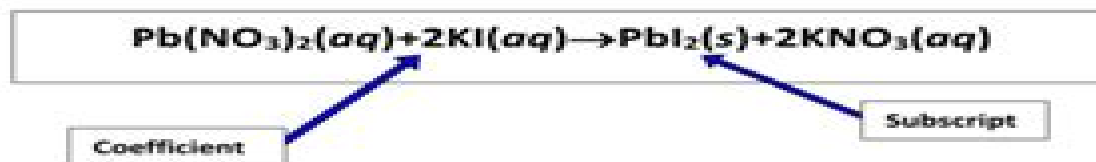
_____ Oxygen



_____ Hydrogen

_____ Oxygen

C. Chemical Equations



1. _____ - # of units of each substance.
2. Individual atom = **atom**
 $2\text{Mg} \Rightarrow 2$ _____ of magnesium
3. Covalent substance = **molecule**
 $3\text{CO}_2 \Rightarrow 3$ _____ of carbon dioxide
4. Ionic substance = **unit**
 $4\text{MgO} \Rightarrow 4$ _____ of magnesium oxide

SYMBOL	MEANING
$\longrightarrow$	produces, forms
+	plus, and
(s)	solid
(l)	liquid
(g)	gas
(aq)	aqueous (solid dissolved in water)
Δ	the reactants are heated

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E Durkheim



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web jul 12 2023 this concept lies at the heart of all calorimetry problems and calculations because the heat released or absorbed at constant pressure is equal to Δh the relationship between heat and Δh_{rxn} is $\Delta h_{\text{rxn}} = q_{\text{rxn}} / \text{mols} \Delta t$ 7 3 14 7 3 14 $\Delta h_{\text{rxn}} = q_{\text{rxn}} / \text{mols} \Delta t$

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web use the information in this chart to answer the following questions 1 if i have 125 grams of water at a temperature of 25.0 c and heat it to 145.0 c what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the process in 1 3 i have 18.7 grams of an unknown liquid chemical at a temperature of 35.0 c

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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