Parts of the Atom

*Fill in the following table about the parts of the atom. (AMU stands for atomic mass unit)

PART OF THE ATOM	ELECTRIC CHARGE	LOCATION IN ATOM	ATOMIC MASS	
proton	proton	Nucleus	2000 X More than an e	
neutron	neutron	Nucleus	2000 X More than an e'	
electron	electron	Orbitals / Energy Levels	1	

*Here's how to predict the structure of any element on the periodic table. For example, look at carbon:

6 C carbon 12.011

Error: Reference source not foundatomic number: This is the number of protons (also the number of electrons)

Error: Reference source not foundchemical symbol: This is an abbreviation for the element

Error: Reference source not foundelement name: This is the name of the element

Error: Reference source not foundatomic mass: This number (rounded off) tells you the number of protons plus the number of neutrons. (Electrons are too small to be included.)

To find the number of protons, look at the atomic number:

Example for carbon: Atomic number: 6

Number of protons: 6

To find the number of neutrons, subtract the atomic number from the atomic mass:

Example for carbon: Atomic mass 12 (rounded) Number of neutrons 12-6=6

To find the number of electrons, look at the atomic number:

Example for carbon: Atomic number: 6 Number of electrons: 6

*Complete the table for the elements with atomic numbers 1-10.

ELEMENT NAME	ATOMIC NUMBER	ATOMIC MASS	PROTONS	NEUTRONS	ELECTRONS
Hydrogen	1	1.00	1	1	1
Helium	2	4.00	2	2	2
Lithium	3	6.91	3	3	3
Beryllium	4	9.01	4	4	4
Boron	5	10.81	5	5	5
Carbon	6	12.01	6	6	6
Nitrogen	7	14.00	7	7	7
Oxygen	8	15.99	8	8	8

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Karin Nielsen-Saines

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Table of Contents Note Taking Worksheet Inside The Atom Answers Packet

- 1. Understanding the eBook Note Taking Worksheet Inside The Atom Answers Packet
 - The Rise of Digital Reading Note Taking Worksheet Inside The Atom Answers Packet
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Note Taking Worksheet Inside The Atom Answers Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Note Taking Worksheet Inside The Atom Answers Packet
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Note Taking Worksheet Inside The Atom Answers Packet
 - Personalized Recommendations
 - Note Taking Worksheet Inside The Atom Answers Packet User Reviews and Ratings
 - Note Taking Worksheet Inside The Atom Answers Packet and Bestseller Lists
- 5. Accessing Note Taking Worksheet Inside The Atom Answers Packet Free and Paid eBooks
 - Note Taking Worksheet Inside The Atom Answers Packet Public Domain eBooks
 - Note Taking Worksheet Inside The Atom Answers Packet eBook Subscription Services
 - Note Taking Worksheet Inside The Atom Answers Packet Budget-Friendly Options

- 6. Navigating Note Taking Worksheet Inside The Atom Answers Packet eBook Formats
 - o ePub, PDF, MOBI, and More
 - Note Taking Worksheet Inside The Atom Answers Packet Compatibility with Devices
 - Note Taking Worksheet Inside The Atom Answers Packet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Note Taking Worksheet Inside The Atom Answers Packet
 - Highlighting and Note-Taking Note Taking Worksheet Inside The Atom Answers Packet
 - Interactive Elements Note Taking Worksheet Inside The Atom Answers Packet
- 8. Staying Engaged with Note Taking Worksheet Inside The Atom Answers Packet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Note Taking Worksheet Inside The Atom Answers Packet
- 9. Balancing eBooks and Physical Books Note Taking Worksheet Inside The Atom Answers Packet
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Note Taking Worksheet Inside The Atom Answers Packet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Note Taking Worksheet Inside The Atom Answers Packet
 - Setting Reading Goals Note Taking Worksheet Inside The Atom Answers Packet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Note Taking Worksheet Inside The Atom Answers Packet
 - Fact-Checking eBook Content of Note Taking Worksheet Inside The Atom Answers Packet
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

• Interactive and Gamified eBooks

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