

A. Identify an equilibrium, the K_{eq} expression for that equilibrium, and the K_{eq} as applied to a particular equilibrium.



Classroom activity: students write the K_{eq} expression for the reaction above. Then, students write the K_{eq} expression for the reaction above.

B. Equilibrium:

• The K_{eq} expression for the reaction above is $K_{eq} = \frac{[C][D]}{[A][B]}$.

• The concentration of each substance for the reaction above is 1.0 M. The concentration of the products is 1.0 M.



$$K_{eq} = \frac{[C][D]}{[A][B]}$$

For K_{eq} in the reaction, the product is formed

1.0	1.0
1.0	1.0
1.0	1.0

Ex. 10: Calculate K_{eq} for the following reaction using the given equilibrium concentrations. Then determine whether the forward or reverse reaction is favored.



$$\frac{[H_2O][CO]}{[H_2O][CO]}$$

$$\frac{[H_2O][CO]}{[H_2O][CO]}$$

$$K_{eq} = 0.0001$$

Note Taking Guide Episode 1202 Answers

G Thomas



Note Taking Guide Episode 1202 Answers:

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Note Taking Guide Episode 1202 Answers Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Note Taking Guide Episode 1202 Answers**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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