

Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

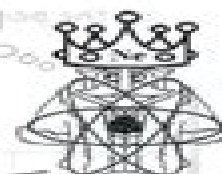
Lithium
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.



Fluorine
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (notorious electron thieves), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.

Francium
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

Iodine
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

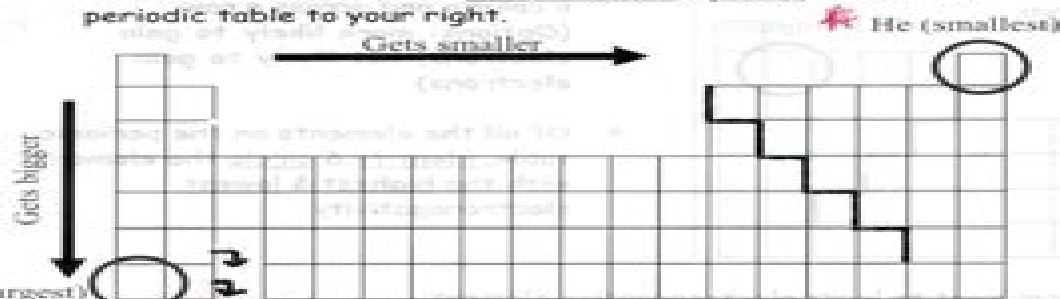


Neon
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

Part 2: Use your textbook to complete the following:

Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

Effective nuclear charge, Z_{eff} , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.

not a good enough answer
4AP Answer

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board of supervisors meeting calendar sandiegocounty gov - Apr 20 2023

as noted on the current board of supervisors meeting calendar a regular meeting of the board is held at 9 a m on tuesday and 9 a m on wednesday in room 310 located at the san diego county administration center 1600 pacific highway san diego california the wednesday session deals with planning and land use matters

board of supervisors alameda county california - Jul 11 2022

agenda board of supervisors meeting tuesday february 28 2023 page 1 board of supervisors regular meeting tuesday february 28 2023 county administration building david haubert vice president district 1 supervisors chambers district 2 1221 oak street lena tam district 3 fifth floor room 512 nate miley president district 4

agenda for the regular meeting of the board of supervisors - Jun 10 2022

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oct 25 2023 matt brown communications specialist publicaffairs sonoma county org 707 565 3040 575 administration drive

suite 104a santa rosa ca 95403 the board of supervisors on tuesday extended american rescue plan act funding for community based organizations that have supported thousands of sonoma county residents through programs

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