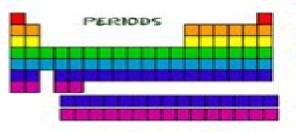
#### Unit 3 Notes: Periodic Table Notes

- John Newlands proposed an organization system based on increasing atomic mass in 1864.
- In 1869 both Lothar Meyer and Dmitri Mendeleev showed a connection between atomic mass and an element's properties.
- Mendeleev published first, and is given credit for this.
- He also noticed a periodic pattern when elements were ordered by increasing <u>Atomic Mass</u>
- By arranging elements in order of increasing atomic mass into columns, Mendeleev created the first Periodic Table.
- This table also predicted the existence and properties of undiscovered elements.
- After many new elements were discovered, it appeared that a number of elements were out of order based on their Properties
- By rearranging the elements based on \_\_\_\_\_\_ Atomic Number \_\_\_\_, the problems with the Periodic Table were corrected.
- This new arrangement creates a periodic repetition of both physical and chemical properties known as the \_\_\_\_\_Periodic Law\_\_\_\_.



Periods are the \_\_\_\_Rows\_\_\_\_

GROUPS

Groups/Families are the Columns

Valence electrons across a period are in the same energy level There are equal numbers of valence electrons in a group.

When elements are arranged in order of increasing <u>Atomic Number</u>,
there is a periodic repetition of their <u>physical</u> and <u>chemical properties</u>

# **Periodic Trends Packet Key**

**K Morrison** 

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