

Unit 3 Notes: Periodic Table Notes

- John Newlands proposed an organization system based on increasing atomic mass in 1864.
- He noticed that both the chemical and physical properties repeated every 8 elements and called this the Law of Octaves.
- In 1869 both Lothar Meyer and Dmitri Mendeleev showed a connection between atomic mass and an element's properties.
- Mendeleev published first, and is given credit for this.
- He also noticed a periodic pattern when elements were ordered by increasing Atomic Mass.
- By arranging elements in order of increasing atomic mass into columns, Mendeleev created the first Periodic Table.
- This table also predicted the existence and properties of undiscovered elements.
- After many new elements were discovered, it appeared that a number of elements were out of order based on their Properties.
- In 1913 Henry Mosley discovered that each element contains a unique number of Protons.
- By rearranging the elements based on Atomic Number, the problems with the Periodic Table were corrected.
- This new arrangement creates a periodic repetition of both physical and chemical properties known as the Periodic Law.

PERIODS

Periods are the Rows

Valence electrons across a period are in the same energy level

GROUPS

Groups/Families are the Columns

There are equal numbers of valence electrons in a group.

- When elements are arranged in order of increasing Atomic Number, there is a periodic repetition of their physical and chemical properties

Periodic Trends Packet Key

K Morrison

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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