

# Solutions Manual for Physical Principles of Plasma Physics and Engineering, 3e by Alexander Fridman, Lawrence Kennedy (All Chapters)

## PART I.

### CHAPTER 2.

**2.7.1. Electron Energy Distribution Functions.** *If electron energy distribution function is Maxwellian and the electron temperature is  $T_e = 1\text{eV}$ , (1) what is the mean velocity  $\langle v \rangle$  of the electrons in this case?, (2) what is the mean electron energy?, (3) which electron energy  $m\langle v \rangle^2/2$  corresponds to the mean velocity?, (4) how can you explain the difference between the two "average electron energies" in terms of standard deviation?*

The mean electron velocity  $\langle v \rangle$  corresponding to the Maxwell velocity distribution function:

$$f(v) = 4\pi n^2 \sqrt{(m/2\pi T_e)^3} \exp(-mv^2/2T_e)$$

can be found after integration as:  $\langle v \rangle = \sqrt{8T_e/\pi m}$ . Numerically at  $T_e = 1\text{eV}$ ,  $\langle v \rangle = 0.7 \cdot 10^8 \text{ cm/sec}$ . The mean electron energy in this case according to (2.1.2) is  $\langle \epsilon \rangle = 1.5 \text{ eV}$ . Electron energy corresponding to the mean velocity can be found as:  $m\langle v \rangle^2/2 = 4T_e/\pi$ , and numerically at  $T_e = 1\text{eV}$  gives the "average" electron energy about 1.3 eV. The difference between these two "average" energies is due to the difference between mean square velocity and square of the mean velocity, which corresponds to the variance or the square of the standard deviation of the velocity distribution.

**2.7.2. Ionization Potentials and Electron Affinities.** *Why are most ionization potentials (Table 2.1.1) greater than electron affinities (Table 2.1.2)?*

# Plasma Physics And Engineering Solution Manual

**Liying Dong**



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