

Percent Yield Tutorial

PRACTICE PROBLEMS:

1. Industrial processes are especially concerned with percent yield. (Why?) For producing phosphoric acid the following reaction is used: $\text{P}_2\text{O}_5(\text{s}) + 3\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{H}_3\text{PO}_4(\text{aq})$. What is the expected mass of H_3PO_4 that will be produced from 50 grams of P_2O_5 reacting in excess water?

$$50 \text{ g } \text{P}_2\text{O}_5 \times \frac{1 \text{ mol}}{142 \text{ g } \text{P}_2\text{O}_5} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ P}_2\text{O}_5} \times \frac{98 \text{ g } \text{H}_3\text{PO}_4}{1 \text{ mol}} = 69.0 \text{ g } \text{H}_3\text{PO}_4$$

expected
⇒ theoretical yield

2. For the above problem, what if 65 grams of H_3PO_4 is actually produced. What is the % yield?

—CrashChemistry—

Percentage Yield Lab Grade 1

M Tight



Percentage Yield Lab Grade 1:

Unveiling the Power of Verbal Art: An Emotional Sojourn through **Percentage Yield Lab Grade 1**

In some sort of inundated with monitors and the cacophony of immediate transmission, the profound power and psychological resonance of verbal artistry often diminish into obscurity, eclipsed by the constant barrage of noise and distractions. Yet, nestled within the lyrical pages of **Percentage Yield Lab Grade 1**, a charming function of literary beauty that impulses with natural thoughts, lies an memorable journey waiting to be embarked upon. Penned with a virtuoso wordsmith, this enchanting opus books viewers on a psychological odyssey, gently exposing the latent potential and profound impact stuck within the intricate web of language. Within the heart-wrenching expanse of the evocative analysis, we can embark upon an introspective exploration of the book is key subjects, dissect its interesting writing style, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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