

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
Calculate: Moles of product based on the limiting reagent. Then convert moles of product to grams. Percent yield = (actual yield / theoretical yield) x 100.

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
 Calculate: Moles of product based on the limiting reagent. Then convert moles of product to grams.



The molar masses of the reactants are: 26.98 g/mol Al, 55.85 g/mol Fe.

Al	26.98 g/mol
Fe	55.85 g/mol
Al ₂ Fe ₃	168.63 g/mol
Fe ₂ O ₃	159.69 g/mol

1. 10.0 g of Al and 10.0 g of Fe are mixed. What is the limiting reagent? What is the theoretical yield of Al₂Fe₃?



2. 10.0 g of Al and 10.0 g of Fe are mixed. What is the limiting reagent? What is the theoretical yield of Al₂Fe₃?

3. 10.0 g of Al and 10.0 g of Fe are mixed. What is the limiting reagent? What is the theoretical yield of Al₂Fe₃?



4. 10.0 g of Al and 10.0 g of Fe are mixed. What is the limiting reagent? What is the theoretical yield of Al₂Fe₃?



Percent Yield And Limiting Reactants Quiz

D Keegan



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