

Periodic and Groups: Valence Electrons: Mastery

Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group										18A	
	1A	2A	3A	4A	5A	6A	7A	8A	9A	10A	11A	12A	13A	14A
1	H	He												
2	Li	Be											B	C
3	Na	Mg											Al	Si
4	K	Ca											Ga	Ge

Thinking like the game "Monopoly", answer that an element is in period 3 and group 1.

What element is in Period 2, and Group 16A? (showing 2 means 2, Group 16A is Oxygen 16A. The element is oxygen.)

What element is in Period 5A and Period 17?

What element is in Group 2A and Period 17?

What group would you find a chlorine in?

Group: _____ Period: _____

What group Magnesium?

Group: _____ Period: _____

Valence Electrons

Valence Electrons are the most important electrons in an atom. Each period contains has the same number of valence electrons. Valence electrons are the electrons that get involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Lithium 3



3 valence electrons

Hydrogen and Helium are both in Group 1A (Hydrogen 1 and Helium 2).

Group 1A, including H and He, all have 1 valence electron. How many valence electrons will they have?

How many electrons?

Group 2A: - Shows you more clearly that there is a 2nd shell of electrons. The elements in Group 2A and 10A are similar to 2 except + He. Helium is not in Group 2 because it has no valence electrons. "He 2 is, 10A" (like Group 1A) has a full outer (2 valence) electron shell. 1A and 10A have full outer shells, just as Helium is in Group 1A.

Periods And Groups Valence Electrons Masses Answers

H Kauffman



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Table of Contents Periods And Groups Valence Electrons Masses Answers

1. Understanding the eBook Periods And Groups Valence Electrons Masses Answers
 - The Rise of Digital Reading Periods And Groups Valence Electrons Masses Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periods And Groups Valence Electrons Masses Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periods And Groups Valence Electrons Masses Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periods And Groups Valence Electrons Masses Answers
 - Personalized Recommendations
 - Periods And Groups Valence Electrons Masses Answers User Reviews and Ratings
 - Periods And Groups Valence Electrons Masses Answers and Bestseller Lists
5. Accessing Periods And Groups Valence Electrons Masses Answers Free and Paid eBooks
 - Periods And Groups Valence Electrons Masses Answers Public Domain eBooks
 - Periods And Groups Valence Electrons Masses Answers eBook Subscription Services
 - Periods And Groups Valence Electrons Masses Answers Budget-Friendly Options

6. Navigating Periods And Groups Valence Electrons Masses Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periods And Groups Valence Electrons Masses Answers Compatibility with Devices
 - Periods And Groups Valence Electrons Masses Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periods And Groups Valence Electrons Masses Answers
 - Highlighting and Note-Taking Periods And Groups Valence Electrons Masses Answers
 - Interactive Elements Periods And Groups Valence Electrons Masses Answers
8. Staying Engaged with Periods And Groups Valence Electrons Masses Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periods And Groups Valence Electrons Masses Answers
9. Balancing eBooks and Physical Books Periods And Groups Valence Electrons Masses Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periods And Groups Valence Electrons Masses Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periods And Groups Valence Electrons Masses Answers
 - Setting Reading Goals Periods And Groups Valence Electrons Masses Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periods And Groups Valence Electrons Masses Answers
 - Fact-Checking eBook Content of Periods And Groups Valence Electrons Masses Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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