

# PERIODIC TABLE

## Atomic Properties of the Elements

**NIST**  
National Institute of  
Standards and Technology  
U.S. Department of Commerce

| Frequently used fundamental physical constants                                                                                                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Physics Laboratory |  | Standard Reference Data |  |  |  |
|---------------------------------------------------------------------------------------------------------------------------------------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--------------------|--|-------------------------|--|--|--|
| For the most accurate values of these and other constants, visit <a href="http://physics.nist.gov/constants">physics.nist.gov/constants</a>             |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | physics.nist.gov   |  | www.nist.gov/srd        |  |  |  |
| 1 second = 9 192 631 770 periods of radiation corresponding to the transition between the two hyperfine levels of the ground state of <sup>133</sup> Cs |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                    |  |                         |  |  |  |
| speed of light in vacuum c = 299 792 458 m s <sup>-1</sup> (exact)                                                                                      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                    |  |                         |  |  |  |
| Planck constant h = 6.626 070 15 × 10 <sup>-34</sup> J s (exact)                                                                                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                    |  |                         |  |  |  |
| elementary charge e = 1.602 176 634 × 10 <sup>-19</sup> C                                                                                               |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                    |  |                         |  |  |  |
| electron mass m <sub>e</sub> = 9.109 383 56 × 10 <sup>-31</sup> kg                                                                                      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                    |  |                         |  |  |  |
| proton mass m <sub>p</sub> = 1.672 621 9 × 10 <sup>-27</sup> kg                                                                                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                    |  |                         |  |  |  |
| fine-structure constant α = 1/137.035 999 074                                                                                                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                    |  |                         |  |  |  |
| Rydberg constant R <sub>∞</sub> = 10 973 731.770 m <sup>-1</sup>                                                                                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                    |  |                         |  |  |  |
| Bohr magneton μ <sub>B</sub> = 9.274 010 078 3 × 10 <sup>-24</sup> J T <sup>-1</sup>                                                                    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                    |  |                         |  |  |  |
| Boltzmann constant k = 1.380 658 × 10 <sup>-23</sup> J K <sup>-1</sup>                                                                                  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                    |  |                         |  |  |  |
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Atomic Number: 58  
Ground-state Ionization Energy (eV): 5.5337

Symbol: **Ce**  
Name: Cerium  
Atomic Weight: 140.116

Ground-state Configuration: [Xe]4f<sup>1</sup>5d<sup>1</sup>6s<sup>2</sup>

\*Based upon <sup>12</sup>C. ( ) indicates the mass number of the longest-lived isotope.

For a description of the data, visit [physics.nist.gov/data](http://physics.nist.gov/data)

NIST SP 966 (September 2010)

# Periodix Table Study Packet

**Karin Nielsen-Saines**



## Periodix Table Study Packet:

**Know Your 'O' Level Chemistry - A Study Guide**, Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement MCAT Study Review Notes - MEGA PACK 900 Pages E Staff, MCAT Study Review Notes MEGA PACK 900 Pages Prepare for the MCAT with this review notes mega pack 900 pages of best review notes from successful MCAT test takers Know all the important facts that you need to succeed on the MCAT From quick facts and mnemonics and everything in between is included in this mega pack Review all the important areas of science Be prepared to ace the test and get admitted into a medical school Content created by highly successful former MCAT test takers with in depth knowledge of what it takes to succeed in this exam ASAP Chemistry: A Quick-Review Study Guide for the AP Exam The Princeton Review, 2019-02-12 Looking for sample exams practice questions and test taking strategies Check out our extended in depth AP chem prep guide Cracking the AP Chemistry Exam LIKE CLASS NOTES ONLY BETTER The Princeton Review s ASAP Chemistry is designed to help you zero in on just the information you need to know to successfully grapple with the AP test No questions no drills just review Advanced Placement exams require students to have a firm grasp of content you can t bluff or even logic your way to a 5 Like a set of class notes borrowed from the smartest student in your grade this book gives you exactly that No tricks or crazy stratagems no sample essays or practice sets Just the facts presented with lots of helpful visuals Inside ASAP Chemistry you ll find Essential concepts terms and functions for AP Chem all explained clearly concisely Diagrams charts and graphs for quick visual reference A three pass icon system designed to help you prioritize learning what you MUST SHOULD and COULD know in the time you have available Ask Yourself questions to help identify areas where you might need extra attention A resource that s perfect for last minute exam prep and for daily class work Topics covered in ASAP Chemistry include Atomic structure Covalent bonding intermolecular forces Thermochemistry Acids bases and more Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while

the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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