

Graphing Periodic Trends Activity

Name: _____ Period: _____

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	<p>Across a period (→) atomic radius tends to _____.</p> <p>Across a period (→) first ionization energy tends to _____.</p> <p>Across a period (→) electronegativity tends to _____.</p>						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	<p>Down a Group (↓) atomic radius tends to _____.</p> <p>Down a Group (↓) first ionization energy tends to _____.</p> <p>Down a Group (↓) electronegativity tends to _____.</p>						
Rb							
Cs							

Periodic Trends Answers Key

Rachel Sandford



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Table of Contents Periodic Trends Answers Key

1. Understanding the eBook Periodic Trends Answers Key
 - The Rise of Digital Reading Periodic Trends Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Answers Key
 - Personalized Recommendations
 - Periodic Trends Answers Key User Reviews and Ratings
 - Periodic Trends Answers Key and Bestseller Lists
5. Accessing Periodic Trends Answers Key Free and Paid eBooks

- Periodic Trends Answers Key Public Domain eBooks
- Periodic Trends Answers Key eBook Subscription Services
- Periodic Trends Answers Key Budget-Friendly Options
- 6. Navigating Periodic Trends Answers Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Answers Key Compatibility with Devices
 - Periodic Trends Answers Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Answers Key
 - Highlighting and Note-Taking Periodic Trends Answers Key
 - Interactive Elements Periodic Trends Answers Key
- 8. Staying Engaged with Periodic Trends Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Answers Key
- 9. Balancing eBooks and Physical Books Periodic Trends Answers Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Answers Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Answers Key
 - Setting Reading Goals Periodic Trends Answers Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Answers Key
 - Fact-Checking eBook Content of Periodic Trends Answers Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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