

PERIODIC TABLE TRENDS

Step 1: Complete the squares for each element by adding the atomic number, name, and atomic mass.

Write the atomic number at the top of the square.

Write the atomic name under the symbol.

Write the atomic mass at the bottom of the square.

Step 2: Determine the number of electrons, protons, and neutrons in each statement.

Step 3: Create a Bohr diagram for each element.

Step 4: Draw a Lewis Dot Structure for each element.

Step 5: Show whether the element is a metal, nonmetal, or metalloid by circling the correct response. Also determine if the element is a solid or gas by circling the correct choice.

Step 6: Use the following colors to shade in the square for each element. You should **ONLY** color in the small square in the upper left-hand corner and not the entire card.

Green – Li and Na

Pink – O and S

Blue – Be and Mg

Purple – F and Cl

Orange – B and Al

Red – C and Si

Brown – N and P

Yellow – He, Ne, and Ar

White – H

Step 7: Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.

1							2
3	4	5	6	7	8	9	10
11	12	13	14	15	16	17	18

Step 8: Put a title on your table.

Step 9: Make a key for each color with the name of the group for that column.

Step 10: Write the column number at the top of each group and the period number on the left side of the table. Be careful, column numbers are not in order from 1-18 and the transition metals, columns 3-12 are not on the table.

Step 11: Draw the red zigzag line that separates the metals from the nonmetals.

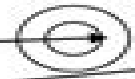
Step 12: Answer the questions on the worksheet using the information on your periodic table.

Be


P = ____

N = ____

E = ____



Bohr Diagram



Dot Structure

Metal

Solid

Nonmetal

Gas

Metalloid

Gas

Periodic Table Basics Packet Answer Key

Tobias Bleicker



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Table of Contents Periodic Table Basics Packet Answer Key

1. Understanding the eBook Periodic Table Basics Packet Answer Key
 - The Rise of Digital Reading Periodic Table Basics Packet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Basics Packet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Basics Packet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Basics Packet Answer Key
 - Personalized Recommendations
 - Periodic Table Basics Packet Answer Key User Reviews and Ratings
 - Periodic Table Basics Packet Answer Key and Bestseller Lists
5. Accessing Periodic Table Basics Packet Answer Key Free and Paid eBooks
 - Periodic Table Basics Packet Answer Key Public Domain eBooks
 - Periodic Table Basics Packet Answer Key eBook Subscription Services
 - Periodic Table Basics Packet Answer Key Budget-Friendly Options
6. Navigating Periodic Table Basics Packet Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Periodic Table Basics Packet Answer Key Compatibility with Devices
- Periodic Table Basics Packet Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Basics Packet Answer Key
 - Highlighting and Note-Taking Periodic Table Basics Packet Answer Key
 - Interactive Elements Periodic Table Basics Packet Answer Key
- 8. Staying Engaged with Periodic Table Basics Packet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Basics Packet Answer Key
- 9. Balancing eBooks and Physical Books Periodic Table Basics Packet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Basics Packet Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Table Basics Packet Answer Key
 - Setting Reading Goals Periodic Table Basics Packet Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Table Basics Packet Answer Key
 - Fact-Checking eBook Content of Periodic Table Basics Packet Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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