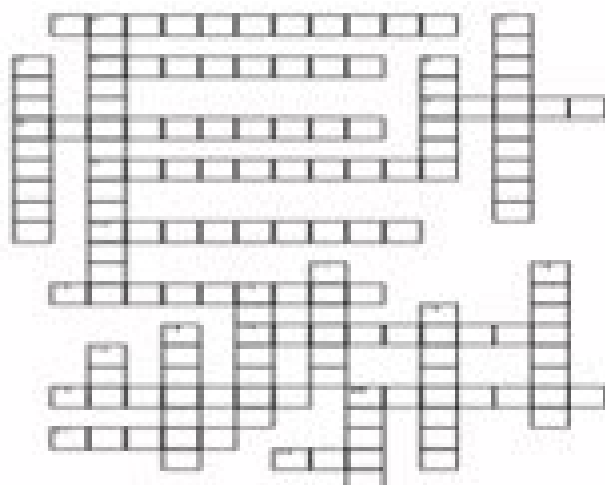


Periodic Table Assignment

Read pages 14 – 19. Complete the crossword puzzle (in pencil) and fill in the period table as instructed below.



ACROSS

1. Synthetic elements with atomic numbers greater than 92
5. This law states that the properties of elements repeat as a function of their atomic number
7. Silicon is in this period
8. Elements were originally placed in order of atomic ____, now they are in order of atomic ____
9. The group that consists of very unreactive gases
10. Number of elements known in 1800
12. Element 101 was named in his honour
17. The periodic table is normally drawn with 18 columns. A more accurate representation has ____ columns
19. Most nonmetals are in this state
20. Nonmetal that is a liquid at SATP
21. The general name given to a column in the periodic table
22. According to figure 6, when aluminum combines with oxygen, three oxygen atoms combine with this many aluminum atoms

DOWN

2. Elements that most closely follow the periodic law
3. The first scientist to organize elements based on repeating patterns
4. The modern name for aka-silicon
6. Most elements are ____, since they fall to the left of the "staircase line"
11. The general name given to a row (left-to-right) in the periodic table
12. Iodine belongs to this group
14. The first alkali metal
15. This element was assigned a relative atomic mass of one because it is the lightest element
16. The state of matter represented by (aq)
18. Silicon is in this group
20. The staircase that separates metals and nonmetals begins below this element

1	H	He																	Fr	Ra																																																																								
2	Li	Be	B	C	N	O	F	Ne									Na	Mg	Al	Si	P	S	Cl	Ar	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Cobalt	Nickel	Cu	Zn	Ga	Ge	As	Se	Br	Kr	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Fr	Ra
3	Na	Mg	Al	Si	P	S	Cl	Ar									K	Ca	Sc	Ti	V	Cr	Mn	Fe	Cobalt	Nickel	Cu	Zn	Ga	Ge	As	Se	Br	Kr	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Fr	Ra								
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Cobalt	Nickel	Cu	Zn	Ga	Ge	As	Se	Br	Kr	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Fr	Ra																								
5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe	Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Fr	Ra																																										
6	Cs	Ba	La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Fr	Ra																																																											
7	Fr	Ra																	Ac	Th	Pa	U	Np	Pu	American	Am	Cm	Bk	Cf	Es	Fm	Mendelevium	No	Lr	Rf	Db	Sg	Bh	Hs	Tl	Pb	Bi	Po	At	Fr	Ra																																														

- 1) Place a large G in boxes that house elements that are gases at SATP, and a large L in boxes that contain elements that are liquids. The rest of the elements are solids; do not write anything in these boxes.
- 2) Darken the "staircase line" that divides metals from non-metals
- 3) Metalloids (i.e. semi-conductors) are elements that have properties midway between metals and non-metals. The metalloids are: B, Si, Ge, As, Sb, Te, Po, At. Identify the metalloids by drawing diagonal lines in these boxes.
- 4) Notice that each group (column) in the periodic table is identified at the top with a number and/or letter. There are two conventions used in the textbook. Label the groups that end in A (e.g. 1A, 11A, 11A, etc).
- 5) Frame and color the following groups (note: H is not an 1A metal): Alkali metals, alkaline earth metals, halogens, noble gases, transition metals, the inner transition elements (made up of lanthanides and actinides).
- 6) Create a legend for your periodic table.

Periodic Table Assignment Answer Key

Clemens Wendtner



Periodic Table Assignment Answer Key:

Assessing Middle and High School Mathematics & Science Sheryn Spencer-Waterman,2013-08-16 For middle and high school teachers of mathematics and science this book is filled with examples of instructional strategies that address students readiness levels interests and learning preferences It shows teachers how to formatively assess their students by addressing differentiated learning targets Included are detailed examples of differentiated formative assessment schedules plus tips on how to collaborate with others to improve assessment processes Teachers will learn how to adjust instruction for the whole class for small groups and for individuals They will also uncover step by step procedures for creating their own lessons infused with opportunities to formatively assess students who participate in differentiated learning activities

College Chemistry I Norman H. Nachtrieb,1960 *Learning Chemistry 7 Solution Book (Year 2023-24)* ,2024-01-02

Learning Elementary Chemistry Class 7 Teacher Resource Book (Academic Year 2023-24) ,2023-05-20 Learning Elementary Chemistry Class 7 Teacher Resource Book Academic Year 2023 24 *Chemistry in the Community (ChemCom)* American Chemical Society,2011-06-17 Touted as the most successful NSF funded project published Chemistry in the Community ChemCom by the American Chemical Society ACS offers a meaningful and memorable chemistry program for all levels of high school students ChemCom covers traditional chemistry topics within the context of societal issues and real world scenarios Centered on decision making activities where students are responsible for generating data in an investigating analyzing that data and then applying their chemistry knowledge to solve the presented problem The text is intensively laboratory based with all 39 of the investigations integrated within the text not separate from the reading With the ChemCom program students learn more organic and biochemistry more environmental and industrial chemistry and more on the particulate nature of matter than other textbooks all within the relevance of solving problems that arise in everyday life Meticulously updated to meet the needs of today s teachers and students the new sixth edition of ChemCom adheres to the new science framework as well as the forthcoming next generation of science standards Incorporating advances in learning and cognitive sciences ChemCom s wide ranging coverage builds upon the concepts and principles found in the National Science Education Standards Correlations are available showing how closely aligned ChemCom is to these and other state standards

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Eye Tracking for STEM Education Research: New Perspectives Pascal Klein, Martin Rusek, Maike Schindler, 2024-04-04 A modern approach to improving education uses the components of experimental scientific research practices based on objective data dissemination of results and the use of modern technologies STEM education research is maturing and new tools and analysis techniques become available As one example eye tracking the recording of persons eye movements has been growing in popularity as it enables researchers to study learning materials effectiveness problem solving and even students approaches during experimentation Eye movements as captured using eye tracking can reveal information about a student s attention and cognition on a process level going well beyond classical product based assessment techniques such as questionnaires or tests

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Science Scope ,2003 **THE Journal** ,1999 **Chemistry in the Laboratory** James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

Study Guide Woolfolk, 2000-06 Biology/science Materials Carolina Biological Supply Company, 1991 *Making Sense of Sensemaking* TJ McKenna, 2025

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College Reading and Study Skills Kathleen T. McWhorter, 1983 Armchair Chemistry David Bradley, Joel Levy, 2018-02-13 Part of the Armchair series Armchair Chemistry is a quick refresher course in how we survey of the science It explains how we evolved from believing in alchemy to discovering modern chemical equations and goes into detail about the lives of the scientists that uncovered them Fascinating and interactive this is ideal for the student brushing up on a subject or for as a clear and accessible companion for beginners and experts alike It contains explanations of different chemical concepts as well as profiles of key scientists and their discoveries It contains clear and concise explanations of different chemical concepts as well as profiles of key scientists and their discoveries A unique feature of the book is its simple step by step exercises Some of these have everyday applications others are theoretical puzzles but all are designed to challenge you and test your newly acquired knowledge The perfect companion for beginners and experts alike Armchair Chemistry does not assume prior knowledge of the subject It conveys the basic elements of chemistry in a way that is clear and accessible no matter your level of ability

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