

Annotated periodic table of the elements

Group names*

- 1 Hydrogen & the alkali metals
- 2 Alkaline earth metals
- 3 Coinage metals
- 12 Volatile metals
- 13 Borogens
- 14 Chalcogens
- 15 Pnictogens
- 16 Chalcogens
- 17 Halogens
- 18 Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

1	2											13	14	15	16	17	18
1 H 1.01 Hydrogen	2 He 4.003 Helium											3 B 10.81 Boron	4 C 12.01 Carbon	5 N 14.01 Nitrogen	6 O 16.00 Oxygen	7 F 18.99 Fluorine	8 Ne 20.18 Neon
3 Li 6.94 Lithium	4 Be 9.012 Beryllium											9 Al 26.98 Aluminum	10 Si 28.09 Silicon	11 P 30.97 Phosphorus	12 S 32.06 Sulfur	13 Cl 35.45 Chlorine	14 Ar 39.95 Argon
5 Na 22.99 Sodium	6 Mg 24.31 Magnesium											15 Ga 69.72 Gallium	16 Ge 72.64 Germanium	17 As 74.92 Arsenic	18 Se 78.96 Selenium	19 Br 79.90 Bromine	20 Kr 83.80 Krypton
7 K 39.10 Potassium	8 Ca 40.08 Calcium											21 In 114.82 Indium	22 Sn 118.71 Tin	23 Sb 121.76 Antimony	24 Te 127.60 Tellurium	25 I 126.90 Iodine	26 Xe 131.29 Xenon
9 Rb 85.47 Rubidium	10 Sr 87.62 Strontium											27 Hg 200.59 Mercury	28 Tl 204.38 Thallium	29 Pb 207.2 Lead	30 Bi 208.98 Bismuth	31 Po 209 Polonium	32 At 210 Astatine
11 Cs 132.91 Cesium	12 Ba 137.33 Barium											33 Au 196.967 Gold	34 Pt 195.084 Platinum	35 Ag 107.868 Silver	36 Cd 112.415 Cadmium	37 In 114.818 Indium	38 Sn 118.710 Tin
13 Fr 223 Francium	14 Ra 226 Radium											39 K 39.0983 Potassium	40 Ca 40.078 Calcium	41 Sc 44.9559 Scandium	42 Ti 47.88 Titanium	43 V 50.9415 Vanadium	44 Cr 51.9961 Chromium
												45 Mn 54.938044 Manganese	46 Fe 55.845 Iron	47 Co 58.933195 Cobalt	48 Ni 58.6934 Nickel	49 Cu 63.546 Copper	50 Zn 65.38 Zinc
												51 Ga 69.723 Gallium	52 Ge 72.64 Germanium	53 As 74.9216 Arsenic	54 Se 78.96 Selenium	55 Br 79.904 Bromine	56 Kr 83.80 Krypton
												57 La 138.905 Lanthanum	58 Ce 140.12 Cerium	59 Pr 140.908 Praseodymium	60 Nd 144.24 Neodymium	61 Pm 144.9126 Promethium	62 Sm 150.36 Samarium
												63 Eu 151.964 Europium	64 Gd 157.25 Gadolinium	65 Tb 158.925 Terbium	66 Dy 162.50 Dysprosium	67 Ho 164.930 Holmium	68 Er 167.26 Erbium
												69 Tm 168.933 Thulium	70 Yb 173.054 Ytterbium	71 Lu 174.967 Lutetium	72 Hf 178.49 Hafnium	73 Ta 180.94788 Tantalum	74 W 183.84 Tungsten
												75 Re 186.207 Rhenium	76 Os 190.23 Osmium	77 Ir 192.222 Iridium	78 Pt 195.084 Platinum	79 Au 196.966569 Gold	80 Hg 200.59 Mercury
												81 Tl 204.38 Thallium	82 Pb 207.2 Lead	83 Bi 208.9804 Bismuth	84 Po 209 Polonium	85 At 210 Astatine	86 Rn 222 Radon
												87 Fr 223 Francium	88 Ra 226 Radium	89 Ac 227 Actinium	90 Th 232.0377 Thorium	91 Pa 231.036888 Protactinium	92 U 238.02891 Uranium
												93 Np 237.048173 Neptunium	94 Pu 244.0642 Plutonium	95 Am 243.061381 Americium	96 Cm 247.070351 Curium	97 Bk 247.070351 Berkelium	98 Cf 251.083288 Californium
												99 Es 252.083288 Einsteinium	100 Fm 257.10 Fermium	101 Md 258.10 Mendelevium	102 No 259.10 Nobelium	103 Lr 260.10 Lawrencium	104 Rg 261.10 Roentgenium
												105 Nh 286.10 Nihonium	106 Fl 289.10 Flerovium	107 Mc 289.10 Moscovium	108 Lv 293.10 Livermorium	109 Ts 294.10 Tennessine	110 Og 294.10 Oganesson

CLASSIFICATION

Metals	Nonmetals
<div>Active</div> <div>Pre-transition</div> <div>Alkali earth</div> <div>Alkaline</div>	<div>Corrosive</div>
Transition	Intermediate
Post-transition	Metalloid
Radioactive	Radioactive

Active metals

- Pre-transition metals**
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Alkali earths**
Closely resemble Ca, Sr and Ba, but many are heavier, and they form mostly pale-colored compounds. Similar to one another, esp. Li to Cu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.
- Actinides**
Soft, dense and reactive metals, those from Th onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

- Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalysts properties. Very weakly electropositive; inertness oxides.
- Frontier metals**
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and tellurates.

Metals

- Metalloids**
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**
H and N are colourless gases, C, P, and Se are metalloids looking solids while Si is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies; low to high electron affinities; moderate to high electronegativity values; relatively poor to moderately strong oxidising agents; tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).
- *as black P, the most stable form in ambient conditions
- Group 3**
This starts as Sc, Ti, and then bifurcates into a La branch i.e. La-Ac, and a Lu branch i.e. Lu-Lr. Trends going down Sc-Lr, and then La branch, resemble those of groups 1–2.
- *Lu branch, more resemble those of groups 1–2

Nonmetals

- Corrosive nonmetals**
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid, I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**
Colorless, colorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Periodic Table Guide Answers

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