

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev stated that certain elements have similar chemical properties
 - (a) elements appeared in regular intervals and subsequent elements were arranged based on that spacing
 - (b) elements
 - (c) atomic numbers
 - (d) mass ratios
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group
 - (b) the physical and chemical properties of elements are functions of atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - (a) adding a row
 - (b) adding a new period
 - (c) adding a new column
 - (d) adding a new group
5. _____ The noble gases that occupy the eighth column are
 - (a) helium
 - (b) neon
 - (c) argon
 - (d) krypton
6. _____ Likewise, the first column in Group 1 has elements number 1, 3, 11, 19, and elements further, group has elements number 11
 - (a) 4
 - (b) 12
 - (c) 18
 - (d) 20
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
 - (a) 11
 - (b) 12
 - (c) 13
 - (d) 14
8. How many elements are in this group?
 - (a) 1
 - (b) 2
 - (c) 3
 - (d) 4
9. What is the atomic number of the element in the same group as element 118?
 - (a) 118
 - (b) 119
 - (c) 120
 - (d) 121

Periodic Law Chapter 5 Test Answers

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Periodic Law Chapter 5 Test Answers:

Decoding **Periodic Law Chapter 5 Test Answers**: Revealing the Captivating Potential of Verbal Expression

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