## Percent, Actual, and Theoretical Yield

LiOH + KCI → LiCI + KOH

a) I began this reaction with 20 grams of lithium hydroxide. What is my theoretical yield of lithium chloride?

- b) I actually produced 6 grams of lithium chloride. What is my percent yield?
- 2) C<sub>3</sub>H<sub>6</sub> + 5 O<sub>2</sub> → 3 CO<sub>2</sub> + 4 H<sub>2</sub>O

a) If I start with 5 grams of C<sub>2</sub>H<sub>6</sub>, what is my theoretical yield of water?

- b) I got a percent yield of 75% How many grams of water did I make?
- Be + 2 HCl → BeCl<sub>2</sub> + H<sub>2</sub>

My theoretical yield of beryllium chloride was 10.7 grams. If my actual yield was 4.5 grams, what was my percent yield?

4) 2 NaCl + CaO → CaCl<sub>2</sub> + Na<sub>2</sub>O

What is my theoretical yield of sodium oxide if I start with 20 grams of calcium oxide?

# **Percent Actual And Theoretical Yield Answer Key**

**M** Tight

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Introduction to General, Organic, and Biochemistry Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15 The most comprehensive book available on the subject Introduction to General Organic and Biochemistry 11th Edition continues its tradition of fostering the development of problem solving skills featuring numerous examples and coverage of current applications Skillfully anticipating areas of difficulty and pacing the material accordingly this readable work provides clear and logical explanations of chemical concepts as well as the right mix of general chemistry organic chemistry and biochemistry An emphasis on real world topics lets readers clearly see how the chemistry will apply to their career

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