

Problem Solving: Obtaining Empirical Formula from Experimental Data

PROCEDURE FOR

Obtaining an Empirical Formula from Experimental Data

1. Write down (or calculate) in grams the masses of each element present in a sample of the compound. If you are given mass percent composition, assume a 100-g sample and calculate the masses of each element from the given percentages.

2. Convert each of the masses in Step 1 to moles by using the appropriate molar mass for each element as a conversion factor.

3. Write down a preliminary formula for the compound using the number of moles of each element from Step 2 as subscripts.

4. Divide all the subscripts in the formula by the smallest subscript.

5. If the subscripts are not whole numbers, multiply all the subscripts by a small whole number (see table) to get whole-number subscripts.

Fractional Subscript	Multiply by This
0.50	2
0.25	4
0.33	3
0.40	5
0.20	2
0.60	5
0.75	4
0.80	5

EXAMPLE 5.15

Obtaining an Empirical Formula from Experimental Data

A laboratory analysis of aspirin determined the following mass percent composition:
C 60.00%
H 4.40%
O 35.52%
Find the empirical formula.

GIVEN: In a 100-g sample, 60.00 g C, 4.40 g H, 35.52 g O.

FOUN: empirical formula

$$60.00 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 5.00 \text{ mol C}$$

$$4.40 \text{ g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 4.34 \text{ mol H}$$

$$35.52 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 2.22 \text{ mol O}$$



The correct empirical formula is $\text{C}_9\text{H}_8\text{O}_4$.

FOR PRACTICE 5.15

Aspirin has the following mass percent composition:

... C 75.00%, H 5.80%, O 19.12%.

What is the empirical formula of aspirin?

Problem Solving Empirical Formula Answer Sheet

Morris Hein, Susan Arena, Cary Willard

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The Practice of Chemistry Donald J. Wink, Sharon Fetzner-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB **Solving Problems in Chemistry** Rod O'Connor, Charles Mickey, Alton Hassell, 1977 **An Introduction to Chemistry** Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

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