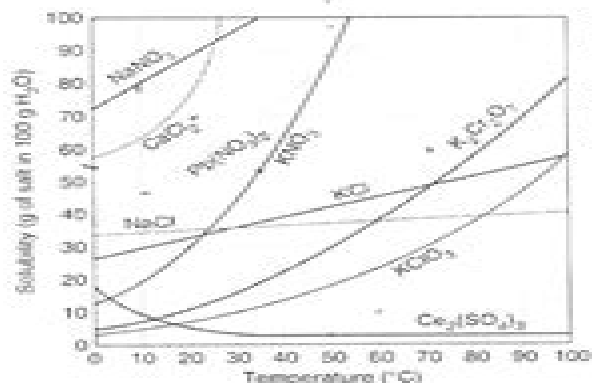


# Worksheet: Solubility Graphs

Name

Key

Use the provided solubility graph to answer the following questions:



For questions 1 - 4 an amount of solute is given, and a temperature is stated. If all of the solute could be dissolved in 100 g of water at the given temperature, would the resulting solution be unsaturated, saturated, or supersaturated?

- 60 g KCl at 70 °C Super Saturated
- 10 g KClO<sub>3</sub> at 60 °C Unsaturated
- 80 g NaNO<sub>3</sub> at 10 °C Saturated
- 70 g CaCl<sub>2</sub> at 20 °C Unsaturated

For questions 5 - 8 a solute and temperature are given. Tell how many grams of each solute must be added to 100 g of water to form a saturated solution at the given temperature.

- Pb(NO<sub>3</sub>)<sub>2</sub> at 10 °C 47g
- Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> at 50 °C 3g
- NaCl at 20 °C 35g
- K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> at 50 °C 30g

For questions 9 and 10 underline the solution that is more concentrated.

- At 10 °C: a saturated solution of KNO<sub>3</sub> or a saturated solution of CaCl<sub>2</sub>
- At 50 °C: a saturated solution of KNO<sub>3</sub> or an unsaturated solution of NaNO<sub>3</sub> consisting of 90 g of the solute dissolved in 100 g of water.

For questions 11 - 12, show your work and circle your final answer.

- If 115 g KNO<sub>3</sub> are added to 100 g of water at 35 °C, how many grams do not dissolve? 115g - 55g = 60g will not dissolve.

- What mass of KCl would be needed to form a saturated solution if the KCl was dissolved in 200 g of water at 80 °C?

$$\frac{52g \text{ KCl}}{100g \text{ H}_2\text{O}} = \frac{104g \text{ KCl}}{200g \text{ H}_2\text{O}}$$

Same Ratio

# Progil Answers Solubility Chemistry

**Joacim Rocklöv**



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