

Chapter 1

Problem Solutions

1.1

- (a) fcc: 8 corner atoms $\times 1/8 = 1$ atom
6 face atoms $\times 1/2 = 3$ atoms
Total of 4 atoms per unit cell

- (b) bcc: 8 corner atoms $\times 1/8 = 1$ atom
1 enclosed atom $= 1$ atom
Total of 2 atoms per unit cell

- (c) Diamond: 8 corner atoms $\times 1/8 = 1$ atom
6 face atoms $\times 1/2 = 3$ atoms
4 enclosed atoms $= 4$ atoms
Total of 8 atoms per unit cell

1.2

- (a) 4 Ga atoms per unit cell

$$\text{Density} = \frac{4}{(5.65 \times 10^{-8})^3} \Rightarrow$$

$$\underline{\text{Density of Ga} = 2.22 \times 10^{22} \text{ cm}^{-3}}$$

4 As atoms per unit cell, so that

$$\underline{\text{Density of As} = 2.22 \times 10^{22} \text{ cm}^{-3}}$$

- (b)

8 Ge atoms per unit cell

$$\text{Density} = \frac{8}{(5.65 \times 10^{-8})^3} \Rightarrow$$

$$\underline{\text{Density of Ge} = 4.44 \times 10^{22} \text{ cm}^{-3}}$$

1.3

- (a) Simple cubic lattice; $a = 2r$

$$\text{Unit cell vol} = a^3 = (2r)^3 = 8r^3$$

$$1 \text{ atom per cell, so atom vol.} = (1) \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{\left(\frac{4\pi r^3}{3} \right)}{8r^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 52.4\%}$$

- (b) Face-centered cubic lattice

$$d = 4r = a\sqrt{2} \Rightarrow a = \frac{d}{\sqrt{2}} = 2\sqrt{2}r$$

$$\text{Unit cell vol} = a^3 = (2\sqrt{2}r)^3 = 16\sqrt{2}r^3$$

$$4 \text{ atoms per cell, so atom vol.} = 4 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{4 \left(\frac{4\pi r^3}{3} \right)}{16\sqrt{2}r^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 74\%}$$

- (c) Body-centered cubic lattice

$$d = 4r = a\sqrt{3} \Rightarrow a = \frac{4}{\sqrt{3}}r$$

$$\text{Unit cell vol.} = a^3 = \left(\frac{4}{\sqrt{3}}r \right)^3$$

$$2 \text{ atoms per cell, so atom vol.} = 2 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{2 \left(\frac{4\pi r^3}{3} \right)}{\left(\frac{4}{\sqrt{3}}r \right)^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 68\%}$$

- (d) Diamond lattice

$$\text{Body diagonal} = d = 8r = a\sqrt{3} \Rightarrow a = \frac{8}{\sqrt{3}}r$$

$$\text{Unit cell vol.} = a^3 = \left(\frac{8}{\sqrt{3}}r \right)^3$$

$$8 \text{ atoms per cell, so atom vol.} = 8 \left(\frac{4\pi r^3}{3} \right)$$

Then

$$\text{Ratio} = \frac{8 \left(\frac{4\pi r^3}{3} \right)}{\left(\frac{8}{\sqrt{3}}r \right)^3} \times 100\% \Rightarrow \underline{\text{Ratio} = 34\%}$$

1.4

From Problem 1.3, percent volume of fcc atoms is 74%. Therefore after coffee is ground,

$$\underline{\text{Volume} = 0.74 \text{ cm}^3}$$

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H Kauffman



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1. Understanding the eBook Problem Physics Manual
 - The Rise of Digital Reading Problem Physics Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Problem Physics Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Problem Physics Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Problem Physics Manual
 - Personalized Recommendations
 - Problem Physics Manual User Reviews and Ratings
 - Problem Physics Manual and Bestseller Lists
5. Accessing Problem Physics Manual Free and Paid eBooks
 - Problem Physics Manual Public Domain eBooks
 - Problem Physics Manual eBook Subscription Services
 - Problem Physics Manual Budget-Friendly Options
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- ePub, PDF, MOBI, and More
- Problem Physics Manual Compatibility with Devices
- Problem Physics Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Problem Physics Manual
 - Highlighting and Note-Taking Problem Physics Manual
 - Interactive Elements Problem Physics Manual
- 8. Staying Engaged with Problem Physics Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Problem Physics Manual
- 9. Balancing eBooks and Physical Books Problem Physics Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Problem Physics Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Problem Physics Manual
 - Setting Reading Goals Problem Physics Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Problem Physics Manual
 - Fact-Checking eBook Content of Problem Physics Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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though different sources cite various types of conclusions all of them serve one of these three primary functions 1

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conclusions should be put to use any time you are writing an essay report or article that proposes or explores an idea issue or event this idea is referred to as a thesis statement and it provides the structure and motivation for the entire piece in other words it answers the why a conclusion on the other see more