

13

ELECTRONS IN ATOMS

Conceptual Curriculum

- ☒ Concrete concepts
- ☐ More abstract concepts or math/problem solving

Standard Curriculum

- ☒ Core content
- ☐ Extension topics

Honors Curriculum

- ☒ Core Honors content
- ☐ Options to accelerate

SECTION 13.1 MODELS OF THE ATOM (pages 361–366)

This section summarizes the development of atomic theory. It also explains the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.

► The Evolution of Atomic Models (pages 361–363)

1. What are the chemical properties of atoms, ions, and molecules related to?

The properties are related to the arrangement of electrons.

2. Complete the table about atomic models and the scientists who developed them.

Scientist	Model of Atom
Dalton	The atom is a solid indivisible mass.
Thomson	The atom is a ball of positive charge with electrons stuck into the ball.
Rutherford	Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space.
Bohr	Electrons are arranged in concentric circular paths around the nucleus.

3. The energy level of an electron is the region around the nucleus where the electron is likely to be moving.
4. Is the following sentence true or false? The electrons in an atom can exist between energy levels. false
5. Circle the letter of the term that completes the sentence correctly. A quantum of energy is the amount of energy required to
- move an electron from its present energy level to the next lower one
 - maintain an electron in its present energy level
 - ☒ move an electron from its present energy level to the next higher one
6. In general, the higher the electron is on the energy ladder, the farther it is from the nucleus.

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