

## **Chapter 7 – Chemical Quantities and Reactions**

### **7.1 THE MOLE**

**Collection Term** // States a specific number of items.

- 1 dozen donuts = 12 donuts
- 1 ream of paper = 500 sheets
- 1 case = 24 cans

**Mole** // A collection that contains the same number of particles as there are carbon atoms in 12.0 g of carbon:  $6.02 \times 10^{23}$  atoms of an element.

**Avogadro's Number** //  $6.02 \times 10^{23}$  atoms of an element.

- 1 mole =  $6.02 \times 10^{23}$  atoms/molecules & this is an equality
- Used to convert moles of a substance to particles
- SIG FIGS count

#### **Subscripts and Moles**

- The subscripts in a formula show...
  - The relationship of atoms in the formula
  - The moles of each element in 1 mole of compound
- The subscripts are used to write conversion factors for moles of each element in 1 mole of a compound
  - 9 moles C/1 mole  $C_9H_8O_4$
  - 8 moles H/1 mole  $C_9H_8O_4$
  - 4 moles O/1 mole  $C_9H_8O_4$

### **7.2-3 MOLAR MASS & CALCULATIONS**

**Molar Mass** // The mass of 1 mole of an element or compound (atomic mass expressed in g).

- For a compound...
  - The sum of the molar masses of the elements in the formula
  - $CaCl_2 = 40.1g/mol (Ca) + (2 \times 35.5g/mol) (Cl)$

#### **Molar Mass Conversion Factors**

- Are fractions written from the molar mass

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**RC Schank**

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