

9.1 Assessment

- The positive charge of a group A metal is equal to its group number. The charge of a group A nonmetal is the group number minus 8.
- from the number of electrons lost
- ite or –ate
- Group 1A metals, 1+; Group 3A (aluminum), 3+; Group 5A nonmetals, 3-
- A monoatomic anion is a single atom with a negative charge; a polyatomic anion is two or more bound atoms with a negative charge.
- a.) K^+ , cation, potassium ion
b.) O^{2-} anion, oxide ion
c.) Sn^{2+} cation, tin (II)
d.) Br^- anion, bromide ion
e.) Be^{2+} cation beryllium ion
f.) Co^{3+} cation, cobalt(III) ion
- a.) NH_4^+
b.) Sn^{2+}
c.) CrO_4^{2-}
d.) NO_3^-

9.2 Assessment

- Write the name of the cation followed by the name of the anion.
- Write the symbol of the cation followed by the symbol of the anion, then use subscripts to balance the charges.
- Write the symbol for the metal ion followed by the formula of the polyatomic ion and balance the charges. Name the cation first followed by the anion.
- a.) $BeCl_2$
b.) Cs_2S
c.) NaI
d.) SrO
- a.) $Cr(NO_3)_3$
b.) $NaClO_4$
c.) $Mg(HCO_3)_2$
d.) $Ca(C_2H_3O_2)_2$
- a.) incorrect; charges are not balanced $MgSO_4$
b.)
c.) incorrect; charges are not balanced $BeCl_2$

9.3 Assessment

- Prefixes indicate the number of atoms of each element in a molecule of the compound.

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**Concepción Jiménez-González, David J.
C. Constable**

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