Name

Date

CHEMICAL BONDING

Answer the following questions.

A chemical bond between atoms results from the attraction between the valence electrons and of different atoms.

((a) muclei

(b) inner electrons

(c) isotopes

(d) Lewis structures

A covalent bond consists of

(a) a shared electron.

(b) a shared electron pair.

(c) two different ions.

(d) an octet of electrons.

If two covalently bonded atoms are identical, the bond is identified as

(a) nonpolar covalent.

(b) polar covalent.

(c) ionic.

(d) dipolar.

A covalent bond in which there is an unequal attraction for the shared electrons is

(a) nonpolar.

(b) polar.

(c) ionic.

(d) dipolar.

Atoms with a strong attraction for electrons they share with another atom exhibit

(a) zero electronegativity.

(b) low electronegativity.

(c) high electronegativity.

(d) Lewis electronegativity.

Bonds that possess between 5% and 50% ionic character are considered to be

(a) ionic.

(b) pure covalent.

(c) polar covalent.

(d) nonpolar covalent.

The greater the electronegativity difference between two atoms bonded together, the greater the bond's percentage of

(a) ionic character.

(b) nonpolar character.

(c) metallic character.

(d) electron sharing.

If electrons involved in bonding spend most of the time closer to one atom rather thanthe other, the bond is polar covalent

If a bond's character is more than 50% ionic, then the bond is called a(n) ionic bond

A bond's character is more than 50% ionic if the electronegativity difference between the two atoms is greater than

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Arturo Cuomo

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