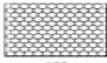
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## Chapter 1

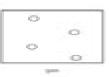
#### Exercise 1.1

1 A solid has a fixed mass and volume / shape. A liquid has a fixed mass but its shape changes to that of the container in which it is placed. A gas has no fixed shape or volume. A gas completely fills the container that it is in.

2



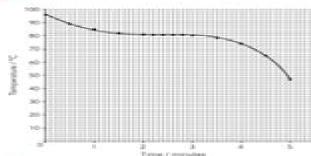




- 3 a H
  - b (
  - A
  - d D
- 4 A freezing (solidification)
  - B melting
    - C condensation
  - D evaporation or boiling
- 5 a radon
  - b radon and nitrogen
  - e nitrogen
  - d cobalt
  - The sample of ethanoic acid is impure.
    The presence of impurities raises the boiling point of a substance.

#### Exercise 1.2





- 7 The student would need to use an oil bath (in place of the water bath) so that the higher temperature could be reached.
- 8 The substance is freezing (solidifying) / turning from liquid to solid.
- The temperature stays constant because energy is being released as the substance solidifies / the molecules are giving out heat as they stop moving from place to place and become organised in a structured lattice arrangement / in the solid the molecules can only vibrate about fixed points / the heat released by the formation of new interactions (forces) between the particles keeps the temperature constant until all the substance is solid.

## **Prentice Hall Chemistry Workbook Answers Chapter 9**

**Roman Wölfel** 

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