

Chapter 12 worksheet 4 (ws12.4) Colligative Properties

A pure substance has well-defined physical properties such as freezing and boiling points. In contrast, **many properties of a solution depend on the concentration of solute**. Properties of solutions that depend on the concentration (**but not on the identity**) of solute are called **colligative** (collective) properties. You will learn how the concentration of a solute affects four properties of a solution. As the concentration of solute increases:

vapor pressure decreases
boiling point increases
freezing point decreases
osmotic pressure increases

We will treat these effects both qualitatively and quantitatively and you will learn why they happen. We will discuss some important applications of these effects.

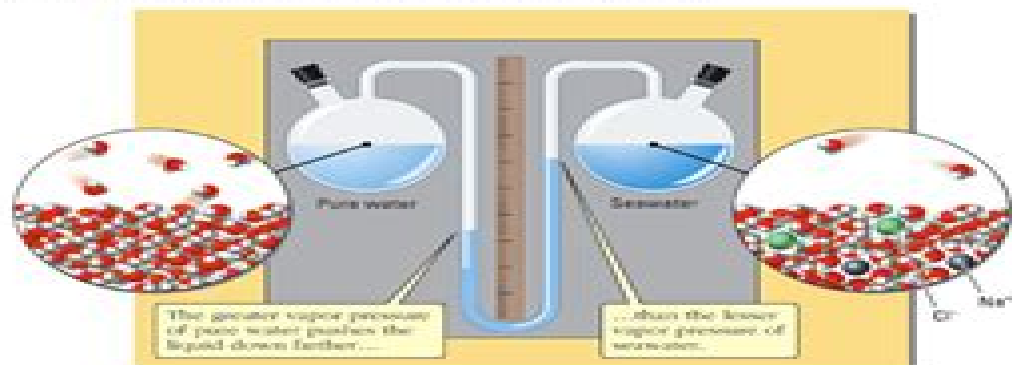
Volatile vs. non-volatile components of a solution

For each colligative property we will usually consider the case of a **non-volatile solute** dissolved in a **volatile solvent** (a typical example is any ionic compound dissolved in water).

A **volatile substance** is one with a measurable vapor pressure at moderate temperatures. In general, volatility increases as intermolecular forces decrease.

Vapor pressure reduction (Raoult's law)

Vapor pressure is the pressure inside a closed container when a liquid is in equilibrium with its own gas phase so the rate of evaporation is equal to the rate of condensation. When a **non-volatile solute** is dissolved in a **volatile solvent**, the vapor pressure of the solvent is reduced. As the concentration of the solute increases, the vapor pressure of the solvent decreases.



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