#### Answers for Predicting Products of Chemical Reactions

For all combustion reactions of hydrocarbons or hydrogen-carbon-oxygen molecules, the products will always be CO: and H:O (assuming a complete reaction).

a. Combustion:	$C_6H_{12}$	+	901	-	6CO <sub>3</sub>	*	6H;O
b. Combustion:	2C <sub>4</sub> H <sub>6</sub>	+	1102	$\rightarrow$	8CO2	+	6H:O
c. Combustion:	$C_0H_{10}O_3$	94	707		6CO <sub>2</sub>		5H <sub>2</sub> O

1. Synthesis: Mg + I₂ → MgI:

Note that Mg can only form Mg2 and I can only form I', so the product will must have a 1:2 cation anion ratio.

2. Double displacement: CuCl₂ + H₂S → CuS + 2HCl Note that the product is **not** H₂Cl₂. It is important to recognize that CuCl₂ is made of three ions, Cu<sup>1+</sup> and two Cl².

Double displacement: NaOH + HCIO<sub>4</sub> → NaCIO<sub>4</sub> + H<sub>2</sub>O
 In this question, you must recognize that perchlorate, ClO<sub>4</sub>, and hydroxide, OH<sup>1</sup>, are polyatomic ions and will not break apart. Also, this is an acid-base reaction, so the products should be salt and water.

Decomposition: ZnCO<sub>2</sub> + heat → ZnO + CO<sub>2</sub>.
 When reactions have heat as a reactant, it is very likely that they will involve decompositions. Carbonate compounds usually decompose to CO<sub>2</sub> and a metal oxide.

5. Single replacement: 2HCl + Zn → ZnCl<sub>2</sub> + H<sub>2</sub>
Note that one reactant is in its elemental form; if a single replacement reaction is going to occur, the species NOT in its elemental form in the reactants (H¹ in this case), will end up in its elemental form in the products (H₂) and the species that IS in its elemental form (Zn) will end up ionized (Zn²). Note that zinc can only form a Zn² ion, so it will have two chloride ions. Note also that hydrogen in its elemental form is H₂, not H.

6. <u>Single replacement</u> 2Na + MgCl<sub>2</sub> → 2NaCl + Mg Again notice that one species is in its elemental form (Na). The magnesium in MgCl<sub>2</sub> is an ion (Mg<sup>2+</sup>), but is transformed into its elemental state (Mg), while the Na is converted into an ion (Na<sup>+</sup>, sodium only forms a +1 ion).

Double replacement CaCl: + K:CO → CaCO: + 2KCl
 Recognize that carbonate is a polyatomic ion (CO₁²) and that the cations are already stable ions (Ca²² and K²).

8. <u>Synthesis</u> 2K + CI: → 2KCI

Note that both materials are elemental species, so the only result could be a synthesis reaction. In the product, K<sup>\*</sup> and CF are formed. Note that we would NOT write the product as K<sub>2</sub>Cl<sub>2</sub>. Ionic compounds are written in their simplest forms.

Double replacement 3BaCl: + 2K.PO₁ → Bas(PO₁): + 6KCl
 Note that phosphate (PO₁) is a polyatomic ion and will not break apart. Since barium is a +2 ion, the barium phosphate will have a 3:2 ratio of Ba:PO₁ in order to balance the charge.

10. <u>Double replacement</u> H:SO<sub>a</sub> + 2KOH → 2H:O + K:SO<sub>a</sub> Notice that sulfate (SO<sub>a</sub><sup>2</sup>) is a polyatomic ion and that potassium exists as a +1 ion (K<sup>2</sup>). Also, this is an acid-base reaction, so the products should be salt and water.

11. Decomposition Al₂(CO₂): + beat → Al₂O₂ + 3CO₂
Another decomposition that will generate CO₂ and a metal oxide (note that Al forms a +3 ion, Al²⁻, and monatomic oxygen will have a -2 charge, O²⁻).

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#### **Table of Contents Predicting Products Hw Packet Answer Key**

- 1. Understanding the eBook Predicting Products Hw Packet Answer Key
  - The Rise of Digital Reading Predicting Products Hw Packet Answer Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Predicting Products Hw Packet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Predicting Products Hw Packet Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Predicting Products Hw Packet Answer Key
  - Personalized Recommendations
  - $\circ\,$  Predicting Products Hw Packet Answer Key User Reviews and Ratings
  - Predicting Products Hw Packet Answer Key and Bestseller Lists
- 5. Accessing Predicting Products Hw Packet Answer Key Free and Paid eBooks
  - Predicting Products Hw Packet Answer Key Public Domain eBooks
  - Predicting Products Hw Packet Answer Key eBook Subscription Services
  - Predicting Products Hw Packet Answer Key Budget-Friendly Options

- 6. Navigating Predicting Products Hw Packet Answer Key eBook Formats
  - o ePub, PDF, MOBI, and More
  - Predicting Products Hw Packet Answer Key Compatibility with Devices
  - Predicting Products Hw Packet Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Predicting Products Hw Packet Answer Key
  - Highlighting and Note-Taking Predicting Products Hw Packet Answer Key
  - Interactive Elements Predicting Products Hw Packet Answer Key
- 8. Staying Engaged with Predicting Products Hw Packet Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Predicting Products Hw Packet Answer Key
- 9. Balancing eBooks and Physical Books Predicting Products Hw Packet Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Predicting Products Hw Packet Answer Key
- 10. Overcoming Reading Challenges
  - o Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Predicting Products Hw Packet Answer Key
  - Setting Reading Goals Predicting Products Hw Packet Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Predicting Products Hw Packet Answer Key
  - Fact-Checking eBook Content of Predicting Products Hw Packet Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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