

First Year Higher Secondary Examination

PART – III

CHEMISTRY

Maximum: 60 Scores

Time: 2 Hours

Cool-off time: 15 Minutes

1. (a) Determine the number of moles present in 0.55 mg of electrons.
 - (i) 1 mole
 - (ii) 2 mole
 - (iii) 1.5 mole
 - (iv) 0.5 mole

(1)
- (b) Give the empirical formula of the following.

(2)

$C_8H_{12}O_6$, C_6H_6 , CH_3COOH , $C_6H_6Cl_6$
- (c) Two elements, carbon and hydrogen combine to form C_2H_6 , C_2H_4 and C_2H_2 , Identify the law illustrated here.

(1)
2. (a) (i) Write the electronic configuration of chromium ($Z = 24$).

(1)
- (ii) Find the number of electrons in the subshells with azimuthal quantum number $l = 2$.

(1)
- (iii) Represent the orbital with quantum numbers $n = 1$ and $l = 0$

(1)
- (b) Give the mathematical representation of Heisenberg's uncertainty principle and its one important significance.

(2)
3. Electron gain enthalpy is one of the important periodic property.
 - (a) Define electron gain enthalpy.

(1)
 - (b) Explain any two factors affecting electron gain enthalpy.

(2)
 - (c) Write the oxidation state and co-valency of Al in $(AlF_6)^{3-}$.

(1)
4. The geometry of the molecule is decided by type of hybridization.
 - (a) Discuss the shape of PCl_5 molecules using hybridization.

(2)
 - (b) Give the reason for the high reactivity of PCl_5 .

(2)
 - (c) Isoelectric species have the same bond order. Among the following, choose the pair having same bond order.

(1)

CN^- , \bar{O}_2 , NO^+ , CN^+

Polytechnic 1st Year Chemistry Question Paper

Michael Seilmaier



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