

Periodic and Groups: Valence Electrons: Mastery

Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group										18A
	1A	2A	3A	4A	5A	6A	7A	8A	9A	10A	11A	12A	13A
1	H	He											
2	Li	Be											
3	Na	Mg											
4	K	Ca											

Thinking like the game "Monopoly", answer that an element is in period 3 and group 1.

What element is in Period 2, and Group 16A? (Hint: 2 means 2, Group 16A is 16th column 16A)

What element is in Period 5 and Group 1?

What element is in Group 2A, and Period 3?

What group number is in Period 4?

What group number is in Period 12A?

Valence Electrons

Valence Electrons are the most important electrons in an atom. Each group contains has the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



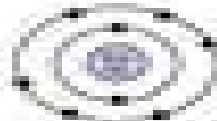
1 valence electron

Helium 2



2 valence electrons

Group 17



7 valence electrons

Hydrogen and Helium are both in Group 1A-2A, but they are both in different groups.

Group 1A, Helium 18A and Group 18A are both in Group 18A. They have the same number of valence electrons. What is it?

What group number is it?

Group 18A: Helium has more shells than Hydrogen, but both are in Group 18A. Helium has 2 valence electrons, Hydrogen has 1. Helium is in Group 18A, Hydrogen is in Group 1A. Helium has 2 valence electrons, Hydrogen has 1. Helium is in Group 18A, Hydrogen is in Group 1A. Helium has 2 valence electrons, Hydrogen has 1. Helium is in Group 18A, Hydrogen is in Group 1A.

Periods And Groups Valence Electrons Masses Answers

Ralph Thomas Myers



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