

# Periodicity - Electronegativity

Name: **Answer Key**

1. Fill in the blanks below, and then draw arrows on to the periodic table indicating the direction of the electronegativity trend.

Generally, electronegativity

**increases**

across periods and

**increases**

up groups.



2. List the elements below in order of increasing electronegativity. (Smallest to largest).

Ba, Ra, Ca, Sr, Mg

**Ra, Ba, Ca, Mg, Sr**

Zn, Ni, Cu, Sc, Ti

**Sc, Ti, Ni, Cu, Zn**

F, B, Cl, Rb, Fr

**Fr, Rb, B, Cl, F**

Nb, Sn, Rb, Fr, I

**Fr, Rb, Nb, Sn, I**

F, Li, Ba, K, Na

**K, Na, Li, Ba, F**

P, Se, Ar, Mg, Ca

**Ca, Mg, Se, P, Ar**

3. Are the elements below in order of decreasing electronegativity? If yes, write "yes". If not, write them in the correct order.

Ca, K, Se, Br

**Br, Se, Ca, K**

I, Br, Cl, F

**F, Cl, Br, I**

Cl, B, Li, Na

**Yes**

Y, Sr, Rb, Fr

**Yes**

At, Sc, Se, Ba

**Yes**

As, Pb, Te, Si

**Si, Te, Pb, As**

4. Briefly explain why electronegativity changes up groups.

Generally, electronegativity increases as you move up in a group. This is because the atomic number decreases, and as a result, there is less distance between the nucleus and valence electrons. In other words, there is a smaller atomic radius.

5. Briefly explain why electronegativity changes across periods.

As you move from left to right, across a period, the nuclear charge of elements increases faster than electron shielding. This results in an increased attraction to electrons.

# Periodicity Worksheet Answer Guide

**Jiyuan Zhang**



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