

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **increases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

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David Baud



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Table of Contents Periodic Table Trends Worksheet Answer Key Bisd303

1. Understanding the eBook Periodic Table Trends Worksheet Answer Key Bisd303
 - The Rise of Digital Reading Periodic Table Trends Worksheet Answer Key Bisd303
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Trends Worksheet Answer Key Bisd303
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Trends Worksheet Answer Key Bisd303
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Trends Worksheet Answer Key Bisd303
 - Personalized Recommendations
 - Periodic Table Trends Worksheet Answer Key Bisd303 User Reviews and Ratings

- Periodic Table Trends Worksheet Answer Key Bisd303 and Bestseller Lists
- 5. Accessing Periodic Table Trends Worksheet Answer Key Bisd303 Free and Paid eBooks
 - Periodic Table Trends Worksheet Answer Key Bisd303 Public Domain eBooks
 - Periodic Table Trends Worksheet Answer Key Bisd303 eBook Subscription Services
 - Periodic Table Trends Worksheet Answer Key Bisd303 Budget-Friendly Options
- 6. Navigating Periodic Table Trends Worksheet Answer Key Bisd303 eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Trends Worksheet Answer Key Bisd303 Compatibility with Devices
 - Periodic Table Trends Worksheet Answer Key Bisd303 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Trends Worksheet Answer Key Bisd303
 - Highlighting and Note-Taking Periodic Table Trends Worksheet Answer Key Bisd303
 - Interactive Elements Periodic Table Trends Worksheet Answer Key Bisd303
- 8. Staying Engaged with Periodic Table Trends Worksheet Answer Key Bisd303
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Trends Worksheet Answer Key Bisd303
- 9. Balancing eBooks and Physical Books Periodic Table Trends Worksheet Answer Key Bisd303
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Trends Worksheet Answer Key Bisd303
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Table Trends Worksheet Answer Key Bisd303
 - Setting Reading Goals Periodic Table Trends Worksheet Answer Key Bisd303
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Table Trends Worksheet Answer Key Bisd303
 - Fact-Checking eBook Content of Periodic Table Trends Worksheet Answer Key Bisd303
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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