

## Periodic Trends Worksheet

*Directions: Use your notes to answer the following questions.*

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to Iodine. This results in smaller size of Fluorine than Iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- |                                          |                  |
|------------------------------------------|------------------|
| a. atomic radius (excluding noble gases) | <b>Decreases</b> |
| b. first ionization energy               | <b>Increases</b> |
| c. electronegativity                     | <b>Increases</b> |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **increases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (Ionization Energy) increases.

# Periodic Lab Table Trend Activities Answer Key

**N Noddings**



## **Periodic Lab Table Trend Activities Answer Key:**

**Hands-On General Science Activities With Real-Life Applications** Pam Walker, Elaine Wood, 2008-04-21 In this second edition of Hands On General Science Activities with Real Life Applications Pam Walker and Elaine Wood have completely revised and updated their must have resource for science teachers of grades 5 12 The book offers a dynamic collection of classroom ready lessons projects and lab activities that encourage students to integrate basic science concepts and skills into everyday life

**The Periodic Table I** D. Michael P. Mingos, 2020-02-05 As 2019 has been declared the International Year of the Periodic Table it is appropriate that Structure and Bonding marks this anniversary with two special volumes In 1869 Dmitri Ivanovitch Mendeleev first proposed his periodic table of the elements He is given the major credit for proposing the conceptual framework used by chemists to systematically inter relate the chemical properties of the elements However the concept of periodicity evolved in distinct stages and was the culmination of work by other chemists over several decades For example Newland s Law of Octaves marked an important step in the evolution of the periodic system since it represented the first clear statement that the properties of the elements repeated after intervals of 8 Mendeleev s predictions demonstrated in an impressive manner how the periodic table could be used to predict the occurrence and properties of new elements Not all of his many predictions proved to be valid but the discovery of scandium gallium and germanium represented sufficient vindication of its utility and they cemented its enduring influence Mendeleev s periodic table was based on the atomic weights of the elements and it was another 50 years before Moseley established that it was the atomic number of the elements that was the fundamental parameter and this led to the prediction of further elements Some have suggested that the periodic table is one of the most fruitful ideas in modern science and that it is comparable to Darwin s theory of evolution by natural selection proposed at approximately the same time There is no doubt that the periodic table occupies a central position in chemistry In its modern form it is reproduced in most undergraduate inorganic textbooks and is present in almost every chemistry lecture room and classroom This first volume provides chemists with an account of the historical development of the Periodic Table and an overview of how the Periodic Table has evolved over the last 150 years It also illustrates how it has guided the research programmes of some distinguished chemists

**Exploring Physical Science in the Laboratory** John T. Salinas , 2019-02-01 This full color manual is designed to satisfy the content needs of either a one or two semester introduction to physical science course populated by nonmajors It provides students with the opportunity to explore and make sense of the world around them to develop their skills and knowledge and to learn to think like scientists The material is written in an accessible way providing clearly written procedures a wide variety of exercises from which instructors can choose and real world examples that keep the content engaging Exploring Physical Science in the Laboratory guides students through the mysteries of the observable world and helps them develop a clear understanding of challenging concepts

*Chemistry: Molecules, Matter, and Change Media Activities Book* Loretta

Jones, Carl Hoeger, Peter William Atkins, Regina Schoenfield-Tacher, 2000-01-15 The Media Activity Book MAB for Jones Atkins Chemistry molecules matter and change contains chapters with lists and descriptions of some of the media available as you study the chapter Each activity begin with a specific textbook reference Then you are given a time estimate of how long it will take to use the media An M media icon in the margin of the textbook means that media exists to support that area of text The media is found in three different places on the website and on two CDs **SourceBook Version 2.1**, 1998 **Build**

**the Brain the Common Core Way** Pamela Nevills, 2014-04-03 Brain compatible learning techniques to help you motivate and challenge your students Boldly take hold of the new Common Core expectations with this fresh innovative resource full of practical tips from international educational expert Pamela Nevills Discover the connection between neuroscience and deep learning that is key to Common Core success Learn to Integrate the standards with current educational practices Understand the science behind Common Core requirements Master teacher lead student centered 21st Century Learning activities Includes school vignettes sample lessons K 12 links and vivid brain imagery Revolutionize your teaching and prepare students for success in school and beyond with this encouraging easy to read guide **Prentice Hall Science Explorer**

Michael J. Padilla, 2002 Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science **The Software Encyclopedia**, 1997

*Chemistry* John S. Phillips, Cheryl Wistrom, 2000 Foundations of Anatomy and Physiology - ePub Ellie Kirov, Alan Needham, 2023-04-01 This new practice manual is designed to provide students with the conceptual foundations of anatomy and physiology as well as the basic critical thinking skills they will need to apply theory to practice in real life settings Written by lecturers Dr Ellie Kirov and Dr Alan Needham who have more than 60 years teaching experience between them the book caters to nursing health science and allied health students at varying levels of understanding and ability Learning activities are scaffolded to enable students to progress to more complex concepts once they have mastered the basics A key advantage of this manual is that it can be used by instructors and students in conjunction with any anatomy and or physiology core textbook or as a standalone resource It can be adapted for learning in all environments including where wet labs are not available Can be used with any other textbook or on its own flexible for teachers and students alike Scaffolded content suitable for students varying learning requirements and available facilities Concept based practical activities can be selected and adapted to align with different units across courses Provides a range of activities to support understanding and

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*Matter, Building Block of the Universe*, 1993 **Matter** Prentice-Hall Staff, 1994 **Learning and Leading with Technology**, 2001 *Science Spectrum* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-03 *ChemCom* American Chemical Society, 1993 *Prentice Hall Chemistry*, 2000

**Chemistry** James E. Brady, John R. Holum, 1995-12-29 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping readers develop skills in analytical thinking and problems solving

**Technical Memorandum** Waterways Experiment Station (U.S.), 1997

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