(2)

(1)

(3) [13]

#### QUESTION 2

2.1 The graph of molecular size versus the boiling point is given below.

The letters A, B and C represent the compounds CH<sub>4</sub>, C<sub>4</sub>H<sub>6</sub> and C<sub>3</sub>H<sub>8</sub> respectively.



- 2.1.1 Define the term boiling point.
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph.
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (A, B or C) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph.
  (2)
- 2.2 Consider the two molecules in the table below.

| NAME<br>OFSUBSTANCE | FORMULA         | MOLACULAR<br>MASS (g) | BOILING POINT |
|---------------------|-----------------|-----------------------|---------------|
| Ammonia             | NH <sub>3</sub> | 17                    | - 33          |
| Phosphine           | PH <sub>3</sub> | ×                     | - 87,4        |

- 2.2.1 Determine the molecular mass (x) of phosphine
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES.

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## **Table of Contents November 2013 Physical Science Question Papers**

- 1. Understanding the eBook November 2013 Physical Science Question Papers
  - The Rise of Digital Reading November 2013 Physical Science Question Papers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying November 2013 Physical Science Question Papers
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an November 2013 Physical Science Question Papers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from November 2013 Physical Science Question Papers
  - Personalized Recommendations
  - November 2013 Physical Science Question Papers User Reviews and Ratings

- November 2013 Physical Science Question Papers and Bestseller Lists
- 5. Accessing November 2013 Physical Science Question Papers Free and Paid eBooks
  - November 2013 Physical Science Question Papers Public Domain eBooks
  - November 2013 Physical Science Question Papers eBook Subscription Services
  - November 2013 Physical Science Question Papers Budget-Friendly Options
- 6. Navigating November 2013 Physical Science Question Papers eBook Formats
  - o ePub, PDF, MOBI, and More
  - November 2013 Physical Science Question Papers Compatibility with Devices
  - November 2013 Physical Science Question Papers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of November 2013 Physical Science Question Papers
  - Highlighting and Note-Taking November 2013 Physical Science Question Papers
  - Interactive Elements November 2013 Physical Science Question Papers
- 8. Staying Engaged with November 2013 Physical Science Question Papers
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers November 2013 Physical Science Question Papers
- 9. Balancing eBooks and Physical Books November 2013 Physical Science Question Papers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection November 2013 Physical Science Question Papers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine November 2013 Physical Science Question Papers
  - Setting Reading Goals November 2013 Physical Science Question Papers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of November 2013 Physical Science Question Papers
  - Fact-Checking eBook Content of November 2013 Physical Science Question Papers
  - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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