

Study Guide – Unit 1 Periodic Table ANSWERS

Structure of an atom

- List the three subatomic particles which make up an atom. Include their location, charge, and atomic mass.

a. Electron	Negative charge	Electron Cloud
b. Proton	Positive charge	Nucleus
c. Neutron	Neutral charge	Nucleus
- In a neutral atom, which two subatomic particles are equal in number to each other?
Protons and Electrons
- What does the atomic number of an element indicate?
The number of protons in the atom of that element
- What is an isotope?
Atoms with the same number of protons, but different masses due to a different number of neutrons
- Will two different atoms of carbon always have the same number of protons?
Yes

Element key

13
Al
26.981
Aluminum

- What are the number of protons, neutrons, and electrons in an atom of aluminum?

a. Protons = 13
b. Neutrons = 14
c. Electrons = 13

Bohr Model

- How do you know how many energy levels (orbitals) a Bohr model will have?
By the period the element is located in
- How many electrons can each energy level hold?
 $1^{\text{st}} = 2$ $2^{\text{nd}} = 8$ $3^{\text{rd}} = 18$

The Periodic Table

- How is the periodic table arranged? Increasing atomic number
- What are the vertical columns called? Groups or Families
- What are the horizontal rows called? Periods
- Why are elements put into groups? Same number of valence electrons
- What is a valence electron? Number of electrons on the outer energy level (shell)
- How many valence electrons each group has
- Which group of metals are highly reactive and why? Alkali Metals. Only 1 val for loss
- Which group of non-metals are highly reactive and why? Halogens. Only 1 val to gain
- Name some physical properties of metals: luster, ductile, malleable, conductor of heat and electricity
- Name some physical properties of non-metals: no luster, brittle, not a good conductor
- What are metalloids and where are they found on the periodic table? Share properties of metals and metalloids. On a zigzag line starting at Boron

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S Ashworth



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