

THERMOCHEMICAL EQUATIONS AND TYPES OF ENTHALPY CHANGES



Thermochemical Equations

A thermochemical equation is a reaction equation accompanied by an enthalpy change (ΔH). In addition to stating the number of moles reactants and the number of moles product, it also displays the amount of heat released or absorbed in the reaction in units of kJ.

Example :

The thermochemical equation is known as:



This means that the formation of 1 mol H_2O hydrogen gas and oxygen gas released energy of 285,85 kJ (exothermic reaction).

Some things that must be considered in the thermochemical equation are

1. THE REACTION COEFFICIENT EXPRESSES MOLES.

Example:



(means that when 1 mol H_2 reacts with 1 mol I_2 to produce 2 mol HI , heat is absorbed by 51.9 kJ).

2. IF THE REACTION IS REVERSED, ΔH CHANGES SIGN

From reaction 1:



3. IF THE REACTIONS ARE MULTIPLIED, THEN ΔH IS MULTIPLIED

From reaction 1:



Notetaking Guide For Thermochemical Equations

Jane D. Smith, George R. Hague



Notetaking Guide For Thermochemical Equations:

Documentation Abstracts ,1980 **Step By Step Guide - How To Balance Chemical Equations (Quick Review Notes)** E Staff, Step By Step Guide How To Balance Chemical Equations Quick Review Notes Learn and review on the go Use Quick Review Chemistry Notes to help you learn or brush up on the subject quickly You can use the review notes as a reference to understand the subject better and improve your grades Easy to remember facts to help you perform better Perfect study notes for all high school and college students 24 Pages **Handbook of Thermochemical Data for Compounds and Aqueous Species** Herbert E. Barner,1994 Thermochemical Reactions D James Benton,2019-06-12 Thermochemistry is the science of analyzing molecular reactions to determine if they are spontaneous energy absorbing or releasing and to predict the product mole ratios and rates Chemical reactions like most other processes tend to follow the path of free energy minimization or entropy maximization This principle forms the mathematical basis for the analytical approach This book is a how to manual filled with many examples and comes with all the code you need to accomplish this task **The Ultimate Chemical Equations Handbook, Teacher Edition** Jane D. Smith,George R. Hague,2011

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